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## CHAPTER 6 SOLUTIONS

6-1 Total carbonate and noncarbonate hardness

Given: Concentrations as ion, see table below

Solution:

a. Begin by converting to mg/L as  $\text{CaCO}_3$

Compound	mg/L as ion	EW/EW	mg/L as $\text{CaCO}_3$
$\text{Ca}^{2+}$	67.2	2.50	168.0
$\text{Mg}^{2+}$	40.0	4.12	164.8
$\text{HCO}_3^-$	--	--	284.0

b. Calculate TH, CH and NCH

$$\text{TH} = 168.0 + 164.8 = 332.8 \text{ mg/L as CaCO}_3$$

$$\text{CH} = \text{HCO}_3^- = 284.0 \text{ mg/L as CaCO}_3$$

$$\text{NCH} = \text{TH} - \text{CH} = 332.8 - 284.0 = 48.8 \text{ mg/L as CaCO}_3$$

6-2 Compute exact TH, CH and NCH

Given: Data from Problem 6-1

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# Introduction To Environmental Engineering 5th Edition Solutions

**Katrin Zwirglmaier**



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