

1. Write equilibrium constant expressions, K_p , for the following reactions.



$$\text{Ans. } K_p = \frac{P_{\text{CO}_2} \times P_{\text{H}_2}}{P_{\text{H}_2\text{O}} \times P_{\text{CO}}}$$



$$\text{Ans. } K_p = \frac{P_{\text{NH}_3}^2}{P_{\text{H}_2}^3 \times P_{\text{N}_2}}$$

2. A mixture containing equal concentrations of methane and steam is passed over a nickel catalyst at 1000 K. The emerging gas has the composition $[\text{CO}] = 0.1027 \text{ mol dm}^{-3}$, $[\text{H}_2] = 0.3080 \text{ mol dm}^{-3}$, and $[\text{CH}_4] = [\text{H}_2\text{O}] = 0.8973 \text{ mol dm}^{-3}$. Assuming this mixture is at equilibrium, calculate the equilibrium constant K_c for the reaction

$$\begin{aligned} \text{Ans. } K_c &= \frac{[\text{CO}][\text{H}_2]^3}{[\text{CH}_4][\text{H}_2\text{O}]} = \frac{0.1027 \text{ mol dm}^{-3} \times (0.3080 \text{ mol dm}^{-3})^3}{0.8973 \text{ mol dm}^{-3} \times 0.8973 \text{ mol dm}^{-3}} \\ &= 3.727 \times 10^{-3} \text{ mol}^2 \text{ dm}^{-6} \end{aligned}$$

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