

**Introduction**

This course is a general chemistry course for students who are interested in pursuing a degree in any of the health-related fields such as nursing, pharmacy, dental hygiene, etc. The course is designed to provide students with a solid foundation in chemistry, which is essential for understanding the biological processes that occur in the human body. The course is divided into two semesters. The first semester covers the topics of atomic structure, chemical bonding, and stoichiometry. The second semester covers the topics of acids and bases, equilibrium, and electrochemistry.

**Prerequisites:** High school chemistry or equivalent. [View the prerequisites page for more information.](#)

**Part 1: Atomic Structure**



1. Define the terms atom, element, and compound. Explain the difference between an atom and a molecule. Describe the structure of an atom, including the nucleus and the electron cloud.
2. Explain the periodic table of elements. Describe the trends in atomic size, ionization energy, and electronegativity.
3. Calculate the atomic weight of an element from its isotopic composition.



4. Calculate the number of protons, neutrons, and electrons in an atom of a given element.
5. Calculate the number of protons, neutrons, and electrons in an ion of a given element.

Element	Atomic Weight	Atomic Number	Number of Protons	Number of Neutrons	Number of Electrons
H	1.008	1	1	1	1
He	4.003	2	2	2	2
Li	6.941	3	3	3	3
Be	9.012	4	4	5	4

6. Calculate the number of protons, neutrons, and electrons in a given ion. Explain the difference between a cation and an anion.

7. Calculate the number of protons, neutrons, and electrons in a given ion. Explain the difference between a cation and an anion.

**Part 2: Chemical Bonding**



8. Define the terms ionic bond, covalent bond, and metallic bond. Explain the difference between an ionic bond and a covalent bond.
9. Explain the octet rule. Describe the trends in ionization energy and electronegativity.
10. Calculate the number of protons, neutrons, and electrons in a given ion. Explain the difference between a cation and an anion.



11. Calculate the number of protons, neutrons, and electrons in a given ion. Explain the difference between a cation and an anion.

Element	Atomic Weight	Atomic Number	Number of Protons	Number of Neutrons	Number of Electrons
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12. Calculate the number of protons, neutrons, and electrons in a given ion. Explain the difference between a cation and an anion.

# Basic Stoichiometry Phet Lab Homework Exercises Answers

**M Walker**



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In some sort of used by monitors and the ceaseless chatter of instant communication, the melodic elegance and psychological symphony created by the prepared term usually disappear in to the back ground, eclipsed by the constant noise and disruptions that permeate our lives. But, nestled within the pages of **Basic Stoichiometry Phet Lab Homework Exercises Answers** an enchanting fictional prize overflowing with natural feelings, lies an immersive symphony waiting to be embraced. Crafted by an outstanding composer of language, that interesting masterpiece conducts viewers on an emotional journey, skillfully unraveling the concealed songs and profound impact resonating within each carefully constructed phrase. Within the depths with this emotional review, we can discover the book is key harmonies, analyze its enthralling publishing style, and submit ourselves to the profound resonance that echoes in the depths of readers souls.

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