

Name _____

Solution Stoichiometry Worksheet

Solve the following solutions Stoichiometry problems:

1. How many grams of silver chromate will precipitate when 150. mL of 0.500 M silver nitrate are added to 100. mL of 0.400 M potassium chromate?



$$\frac{0.150 \text{ L AgNO}_3}{1 \text{ L}} \times \frac{0.500 \text{ moles AgNO}_3}{1 \text{ L}} \times \frac{1 \text{ moles Ag}_2\text{CrO}_4}{2 \text{ moles AgNO}_3} \times \frac{331.74 \text{ g Ag}_2\text{CrO}_4}{1 \text{ moles Ag}_2\text{CrO}_4} = 12.4 \text{ g Ag}_2\text{CrO}_4$$

$$\frac{0.100 \text{ L K}_2\text{CrO}_4}{1 \text{ L}} \times \frac{0.400 \text{ moles K}_2\text{CrO}_4}{1 \text{ L}} \times \frac{1 \text{ moles Ag}_2\text{CrO}_4}{1 \text{ moles K}_2\text{CrO}_4} \times \frac{331.74 \text{ g Ag}_2\text{CrO}_4}{1 \text{ moles Ag}_2\text{CrO}_4} = 13.3 \text{ g Ag}_2\text{CrO}_4$$

2. How many mL of 0.280 M barium nitrate are required to precipitate as barium sulfate all the sulfate ions from 25.0 mL of 0.350 M aluminum sulfate? (93.8 mL barium nitrate)



$$\frac{0.0250 \text{ L Al}_2(\text{SO}_4)_3}{1 \text{ L}} \times \frac{0.350 \text{ moles Al}_2(\text{SO}_4)_3}{1 \text{ L}} \times \frac{3 \text{ moles Ba(NO}_3)_2}{1 \text{ moles Al}_2(\text{SO}_4)_3} \times \frac{1 \text{ L}}{0.280 \text{ moles Ba(NO}_3)_2} = 0.0938 \text{ L Ba(NO}_3)_2$$

3. 25.0 mL of 0.350 M NaOH are added to 45.0 mL of 0.125 M copper (II) sulfate. How many grams of copper (II) hydroxide will precipitate?



$$\frac{0.0250 \text{ L NaOH}}{1 \text{ L NaOH}} \times \frac{0.350 \text{ moles NaOH}}{1 \text{ L NaOH}} \times \frac{1 \text{ moles Cu(OH)}_2}{2 \text{ moles NaOH}} \times \frac{97.57 \text{ g Cu(OH)}_2}{1 \text{ mole Cu(OH)}_2} = 0.427 \text{ g Cu(OH)}_2$$

$$\frac{0.0450 \text{ L CuSO}_4}{1 \text{ L NaOH}} \times \frac{0.125 \text{ moles CuSO}_4}{1 \text{ L NaOH}} \times \frac{1 \text{ moles Cu(OH)}_2}{1 \text{ moles CuSO}_4} \times \frac{97.57 \text{ g Cu(OH)}_2}{1 \text{ mole Cu(OH)}_2} = 0.549 \text{ g Cu(OH)}_2$$

4. What volume of 0.415 M silver nitrate will be required to precipitate as silver bromide all the bromide ion in 35.0 mL of 0.128 M calcium bromide?



$$\frac{0.0350 \text{ L CaBr}_2}{1 \text{ L CaBr}_2} \times \frac{0.128 \text{ moles CaBr}_2}{1 \text{ L CaBr}_2} \times \frac{2 \text{ moles AgNO}_3}{1 \text{ moles CaBr}_2} \times \frac{1 \text{ L AgNO}_3}{0.415 \text{ mole AgNO}_3} = 0.0216 \text{ L AgNO}_3$$

5. What volume of 0.496 M HCl is required to neutralize 20.0 mL of 0.809 M sodium hydroxide?



$$\frac{0.0200 \text{ L NaOH}}{1 \text{ L NaOH}} \times \frac{0.809 \text{ mole NaOH}}{1 \text{ L NaOH}} \times \frac{1 \text{ mole HCl}}{1 \text{ mole NaOH}} \times \frac{1 \text{ L HCl}}{0.496 \text{ mole HCl}} = 0.0326 \text{ L HCl}$$

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AN Whitehead



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Table of Contents Solution Stoichiometry Worksheet Answer Key

1. Understanding the eBook Solution Stoichiometry Worksheet Answer Key
 - The Rise of Digital Reading Solution Stoichiometry Worksheet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Solution Stoichiometry Worksheet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solution Stoichiometry Worksheet Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solution Stoichiometry Worksheet Answer Key
 - Personalized Recommendations
 - Solution Stoichiometry Worksheet Answer Key User Reviews and Ratings
 - Solution Stoichiometry Worksheet Answer Key and Bestseller Lists

5. Accessing Solution Stoichiometry Worksheet Answer Key Free and Paid eBooks
 - Solution Stoichiometry Worksheet Answer Key Public Domain eBooks
 - Solution Stoichiometry Worksheet Answer Key eBook Subscription Services
 - Solution Stoichiometry Worksheet Answer Key Budget-Friendly Options
6. Navigating Solution Stoichiometry Worksheet Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Solution Stoichiometry Worksheet Answer Key Compatibility with Devices
 - Solution Stoichiometry Worksheet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solution Stoichiometry Worksheet Answer Key
 - Highlighting and Note-Taking Solution Stoichiometry Worksheet Answer Key
 - Interactive Elements Solution Stoichiometry Worksheet Answer Key
8. Staying Engaged with Solution Stoichiometry Worksheet Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solution Stoichiometry Worksheet Answer Key
9. Balancing eBooks and Physical Books Solution Stoichiometry Worksheet Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solution Stoichiometry Worksheet Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Solution Stoichiometry Worksheet Answer Key
 - Setting Reading Goals Solution Stoichiometry Worksheet Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Solution Stoichiometry Worksheet Answer Key
 - Fact-Checking eBook Content of Solution Stoichiometry Worksheet Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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