

Physical Science Chapter 6 Test Review Sheet 2014-15

Assessments- This is the portion of the test which you, the student, will be permitted to retest on.

On the assessment, students will be required to:

1. Write out the names and formulas for a binary ionic compound.
2. Write out the names and formulas for a covalent (molecular) compound.

Students will know the following prefixes used for naming covalent/molecular compounds:

Mono-	1	Di-	2	Tri-	3
Tetra-	4	Penta-	5	Hexa-	6
Hepta-	7	Octa-	8	Nona-	9
Deca-	10				

Multiple Choice/Short Answer/Essay

1. Students will know, understand, and be able to apply the following terms:

electron dot diagram	ion	anion	cation
chemical bond	ionic bond	chemical formula	
covalent bond	molecule	polar covalent bond	metallic bond
polyatomic ion	semiconductor	molecular formula	empirical formula
hydrogen bond	formula unit	nonpolar covalent bond	

2. Students will know the composition, physical and chemical properties of ionic and covalent compounds.
3. Students will be able to predict and explain why elements lose or gain electrons to form ions.
4. Students will describe which groups in the periodic table are most willing to lose or gain electrons and explain why.
5. Students will know the different diatomic molecules as described in class (incomplete list in textbook)
6. Students will know what factors play a role in whether a molecular compound will be polar or nonpolar.

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Chao Zhang



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