

Introduction

This course is a general introduction to the field of chemistry. It is designed to give students a solid foundation in the basic principles of chemistry, including atomic structure, chemical bonding, and stoichiometry. The course is intended for students who are interested in pursuing a degree in chemistry or a related field. The course is divided into two semesters. The first semester covers the topics of atomic structure, chemical bonding, and stoichiometry. The second semester covers the topics of chemical equilibrium, acid-base chemistry, and electrochemistry.

Prerequisites: High school chemistry (or equivalent) and high school physics (or equivalent).
[View the course description for CHEM 101 at the University of North Carolina at Chapel Hill.](#)

Part 1: Atomic Structure



1. Explain the structure of an atom, including the nucleus and the electron cloud. Describe the properties of the nucleus, including its mass and charge. Explain the role of the nucleus in determining the identity of an element.
2. Describe the structure of the periodic table, including the groups and periods. Explain the periodic trends in atomic size, ionization energy, and electronegativity.



3. Explain the structure of the periodic table, including the groups and periods. Explain the periodic trends in atomic size, ionization energy, and electronegativity.

Element	Atomic Number	Group Number	Period Number	Block	Classification
H	1	1	1	s	Alkali metal
He	2	18	1	p	Noble gas
Li	3	1	2	s	Alkali metal
Be	4	2	2	s	Alkaline earth metal
B	5	13	2	p	Metalloid
C	6	14	2	p	Nonmetal
N	7	15	2	p	Nonmetal
O	8	16	2	p	Nonmetal
F	9	17	2	p	Halogen
Ne	10	18	2	p	Noble gas

4. Explain the structure of the periodic table, including the groups and periods. Explain the periodic trends in atomic size, ionization energy, and electronegativity.

Part 2: Chemical Bonding



5. Explain the structure of the periodic table, including the groups and periods. Explain the periodic trends in atomic size, ionization energy, and electronegativity.
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10. Explain the structure of the periodic table, including the groups and periods. Explain the periodic trends in atomic size, ionization energy, and electronegativity.

Element	Atomic Number	Group Number	Period Number	Block	Classification
H	1	1	1	s	Alkali metal
He	2	18	1	p	Noble gas
Li	3	1	2	s	Alkali metal
Be	4	2	2	s	Alkaline earth metal
B	5	13	2	p	Metalloid
C	6	14	2	p	Nonmetal
N	7	15	2	p	Nonmetal
O	8	16	2	p	Nonmetal
F	9	17	2	p	Halogen
Ne	10	18	2	p	Noble gas

11. Explain the structure of the periodic table, including the groups and periods. Explain the periodic trends in atomic size, ionization energy, and electronegativity.

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M Tight



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