

Biology Exam #2 (Chapters 2, 3, 4, 6, 7, 8)

Chapter 2:

Atoms are made up of electrons (- outside the nucleus), protons(+inside the nucleus), and neutrons (0 inside the nucleus)

Atomic number= number of protons

Mass number=protons + nucleus

Matter is composed of elements and exists as solid, liquid, gas, or plasma

Element: a substance that cannot be broken down by chemical means to a simpler substance that has different physical properties.

Basic elements to life include C, H, N, O, P, S

Life's diversity is based on carbon, almost all the molecules of a cell are composed of carbon

Organic molecules: carbon-based molecules, molecules have C and H, those with only C and H are called hydrocarbons

Isotopes: atoms of an element with different number of neutrons

Radioactive isotope emits energy as rays and subatomic particles

Uses of radiation include medical imaging, killing bacteria and fungi, killing cancer cells

For neutral atoms, the number of valence electrons is equal to the atom's main group number

ATP: currency of the cell, high energy molecule, undergoes hydrolysis which releases energy

Polar covalent bonds: sharing between two atoms that is unequal

Oxygen is more electronegative than Hydrogen

Second strongest

Nonpolar covalent bonds: sharing of electrons between two atoms is fairly equal

Strongest

Carbon bonds: four electrons on outer shell

Electronegativity: attraction of an atom for electrons in covalent bond

Hydrogen bonding: attraction between negative oxygen and positive hydrogen

Bonds is weak individually, but strong together(weakest)

Biology Study Answers

Heather Wilson Ashworth



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dicarbon hexahydride oxygen carbon dioxide water synthesis equation copper sulfur copper i sulfide decomposition equation copper i sulfide copper sulfur single displacement equation iron lead ii nitrate iron iii nitrate