

VIVA QUESTIONS WITH ANSWERS

Q. 1. What is chemical kinetics ?

Ans. Chemical kinetics is the branch of chemistry which deals with the study of the rates of reactions and their mechanisms.

Q. 2. What do you understand by the rate of reaction ?

Ans. The rate of a reaction is defined as the change in the molar concentration of any one of the reactant or the product per unit time.

Q. 3. What are the units of the rate of reaction ?

Ans. $\text{mol L}^{-1} \text{s}^{-1}$ (moles per litre per second).

Q. 4. What are the factors on which the rate of a reaction depends ?

Ans. The rate of reaction depends upon : (i) Nature of the reactants ; (ii) Concentration of the reactants ; (iii) Temperature ; (iv) Presence of catalyst ; and (v) Presence of radiations.

Q. 5. What is the law of mass action ?

Ans. Law of mass action states that the rate of a reaction is directly proportional to the product of molar concentrations of the reactants.

Q. 6. What is temperature coefficient of a reaction ?

Ans. Temperature coefficient of a reaction is the ratio of rate constants at two temperatures differing by 10° .

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Q. 7. What are the units of rate constant for zero order reactions ?

Ans. Same as the rate of reaction, i.e. $\text{mol L}^{-1} \text{s}^{-1}$.

Q. 8. What are the units of rate constant for first order reactions ?

Ans. s^{-1} .

Q. 9. What is the effect of temperature on the rate constant of a reaction ?

Ans. It generally increases with increase in temperature.

Q. 10. Why certain reactions are very fast ?

Ans. Because they have very low activation energy.

Q. 11. What is threshold energy ?

Ans. It is the minimum energy which the colliding molecules must possess so as to have effective collision.

Q. 12. Why reactions with molecularity more than three are rare ?

Ans. Because simultaneous collision between more than three particles is rare on the basis of probability considerations.

Q. 13. "For an exothermic reaction activation energy for the forward reaction is less than that for the reverse reaction." Justify.

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L Darling-Hammond



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