

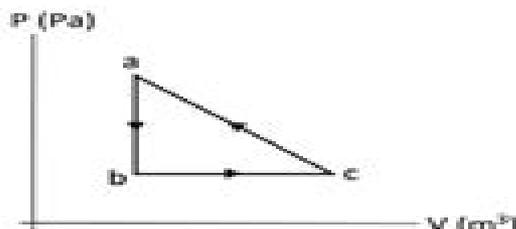
First Law of Thermodynamics

Grade

Name _____ Date _____

Use the following P-V diagram to answer questions about this thermodynamic cycle consisting of three processes labeled ab , bc , and ca .

$P_a = 3.00 \times 10^5 \text{ Pa}$, $P_b = 1.00 \times 10^5 \text{ Pa}$, $V_a = 3 \times 10^{-3} \text{ m}^3$, $V_c = 9 \times 10^{-3} \text{ m}^3$.



a) Identify the type of thermodynamic processes for each part of the cycle (ab , bc , ca) as:

- | | | | | | |
|--------|--------------|--------------|----------------|----------------|------------------|
| ab : | i) adiabatic | ii) isobaric | iii) isochoric | iv) isothermal | v) none of these |
| bc : | i) adiabatic | ii) isobaric | iii) isochoric | iv) isothermal | v) none of these |
| ca : | i) adiabatic | ii) isobaric | iii) isochoric | iv) isothermal | v) none of these |

b) For the process ab , determine (i) the work W_{ab} done by the gas and (ii) the change in internal energy of the system for a heat loss of 900 J. [Hints: for (i) recall that W is the area under the curve; for (ii) use the 1st Law.]

c) For the process bc , determine (i) the work W_{bc} done by the gas and (ii) the change in internal energy of the system for a heat gain of 1250 J.

d) For the process ca , determine (i) the work W_{ca} done by the gas and (ii) the change in internal energy of the system for a heat gain of 150 J.

e) Determine the total work W_{tot} done by the gas in the cycle. Compare this to the area of the triangle. Note that both are negative. Also note that the total change in internal energy is zero, as it must be when the system returns to the same state.

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