

Worksheet 2: Binary Ionic Compounds (variable charge metals) - metals from groups 3 - 12 and 14 or 4A (Pb and Sn) - charge of anion is constant - use charge of anion to figure out charge of metal

1. Name the following ionic compounds:

a. SnO_2	oxide = O^{-2}	Tin (IV) oxide
b. Mn_2O_7	oxide = O^{-2}	Manganese (VII) oxide
c. FeN	nitride = N^{-3}	Iron (III) nitride
d. Cu_3N_2	nitride = N^{-3}	Copper (II) nitride
e. TiF_3	fluoride = F^{-1}	Titanium (III) fluoride
f. Cu_2S	sulfide = S^{-2}	Copper (I) sulfide
g. Fe_2S_3	sulfide = S^{-2}	Iron (III) sulfide
h. CuBr	bromide = Br^{-1}	Copper (I) bromide
i. Co_3N_2	nitride = N^{-3}	Cobalt (II) nitride
j. CoF_2	fluoride = F^{-1}	Cobalt (II) fluoride

2. Write the formula for the following ionic compounds:

a. Copper (I) nitride	Cu_3N
b. Cobalt (I) fluoride	CoF
c. Titanium (IV) oxide	TiO_2
d. Iron (II) nitride	Fe_3N_2
e. Iron (III) chloride	FeCl_3
f. Copper (II) oxide	CuO
g. Rhodium (II) oxide	RhO
h. Tin (IV) sulfide	SnS_2
i. Manganese (IV) nitride	Mn_3N_4
j. Copper (I) fluoride	CuF
k. Cobalt (II) sulfide	CoS
l. Iron (III) oxide	Fe_3O_2

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J Elliott



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Table of Contents Writing And Naming Binary Compounds Worksheet Answer Key

1. Understanding the eBook Writing And Naming Binary Compounds Worksheet Answer Key
 - The Rise of Digital Reading Writing And Naming Binary Compounds Worksheet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Writing And Naming Binary Compounds Worksheet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Writing And Naming Binary Compounds Worksheet Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Writing And Naming Binary Compounds Worksheet Answer Key
 - Personalized Recommendations
 - Writing And Naming Binary Compounds Worksheet Answer Key User Reviews and Ratings
 - Writing And Naming Binary Compounds Worksheet Answer Key and Bestseller Lists

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 - Writing And Naming Binary Compounds Worksheet Answer Key eBook Subscription Services
 - Writing And Naming Binary Compounds Worksheet Answer Key Budget-Friendly Options
6. Navigating Writing And Naming Binary Compounds Worksheet Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Writing And Naming Binary Compounds Worksheet Answer Key Compatibility with Devices
 - Writing And Naming Binary Compounds Worksheet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Writing And Naming Binary Compounds Worksheet Answer Key
 - Highlighting and Note-Taking Writing And Naming Binary Compounds Worksheet Answer Key
 - Interactive Elements Writing And Naming Binary Compounds Worksheet Answer Key
8. Staying Engaged with Writing And Naming Binary Compounds Worksheet Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Writing And Naming Binary Compounds Worksheet Answer Key
9. Balancing eBooks and Physical Books Writing And Naming Binary Compounds Worksheet Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Writing And Naming Binary Compounds Worksheet Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Writing And Naming Binary Compounds Worksheet Answer Key
 - Setting Reading Goals Writing And Naming Binary Compounds Worksheet Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Writing And Naming Binary Compounds Worksheet Answer Key
 - Fact-Checking eBook Content of Writing And Naming Binary Compounds Worksheet Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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