

Chemical Equilibrium

Introduction

Chemical equilibrium is a state of dynamic balance between the forward and reverse reactions. It is reached when the rates of the forward and reverse reactions are equal. At equilibrium, the concentrations of the reactants and products remain constant over time. The equilibrium constant, K_c , is a measure of the extent of the reaction. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient. The equilibrium constant is a function of temperature only.

Equilibrium Constant (K_c) is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient. The equilibrium constant is a function of temperature only.

Part 1: Equilibrium Constant



- Write the balanced chemical equation for the reaction between hydrogen gas and iodine gas to form hydrogen iodide gas. The equilibrium constant for this reaction is $K_c = 50.0$ at 425°C.
- Write the expression for the equilibrium constant, K_c , for the reaction between hydrogen gas and iodine gas to form hydrogen iodide gas.
- Calculate the equilibrium constant, K_c , for the reaction between hydrogen gas and iodine gas to form hydrogen iodide gas, given the following concentrations at equilibrium: $[H_2] = 0.10 \text{ M}$, $[I_2] = 0.10 \text{ M}$, and $[HI] = 1.0 \text{ M}$.



14. Complete the following table for the reaction between hydrogen gas and iodine gas to form hydrogen iodide gas.

Species	Initial Concentration (M)	Change in Concentration (M)	Equilibrium Concentration (M)	Equilibrium Constant (K_c)	Equilibrium Concentration (M)
H_2	1.000	-0.500	0.500	50.0	0.500
I_2	1.000	-0.500	0.500	50.0	0.500
HI	0.000	+1.000	1.000	50.0	1.000
HI	1.000	-0.500	0.500	50.0	0.500

15. Calculate the equilibrium constant, K_c , for the reaction between hydrogen gas and iodine gas to form hydrogen iodide gas, given the following concentrations at equilibrium: $[H_2] = 0.10 \text{ M}$, $[I_2] = 0.10 \text{ M}$, and $[HI] = 1.0 \text{ M}$.

Part 2: Equilibrium Constant



- Write the balanced chemical equation for the reaction between hydrogen gas and iodine gas to form hydrogen iodide gas. The equilibrium constant for this reaction is $K_c = 50.0$ at 425°C.
- Write the expression for the equilibrium constant, K_c , for the reaction between hydrogen gas and iodine gas to form hydrogen iodide gas.

$$K_c = \frac{[HI]^2}{[H_2][I_2]}$$

16. Complete the following table for the reaction between hydrogen gas and iodine gas to form hydrogen iodide gas.

Species	Initial Concentration (M)	Change in Concentration (M)	Equilibrium Concentration (M)	Equilibrium Constant (K_c)	Equilibrium Concentration (M)
H_2	1.000	-0.500	0.500	50.0	0.500
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HI	0.000	+1.000	1.000	50.0	1.000
HI	1.000	-0.500	0.500	50.0	0.500

17. Write the balanced chemical equation for the reaction between hydrogen gas and iodine gas to form hydrogen iodide gas.

Basic Stoichiometry Phet Lab Answers

Jiyuan Zhang



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