

Name Key

1. Hydrogen, H₂, is used as rocket fuel. The hydrogen is burned in oxygen, O₂, according to the following equation. What is the enthalpy change if 200.0 g of hydrogen, H₂ is burned?



$$\frac{200.0 \text{ g H}_2}{2.016 \text{ g H}_2} \times \frac{1 \text{ mol H}_2}{1} = 99.2 \text{ mol H}_2$$

$$99.2 \text{ mol H}_2 \times \frac{-484 \text{ kJ}}{2 \text{ mol H}_2} = -24,006 \text{ kJ}$$

-24,006 kJ

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



What is ΔH for the following reaction?

reverse so change sign of ΔH - that's in 1/2 quantity

$$\frac{1}{2} \text{CO}_2(\text{g}) + \text{SO}_2(\text{g}) \rightarrow \frac{1}{2} \text{CS}_2(\text{l}) + \frac{3}{2} \text{O}_2(\text{g})$$

$$+ \frac{1077 \text{ kJ}}{2} = 538.5 \text{ kJ}$$

3. Following is the equation for the combustion of acetone:



How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

$$\frac{-853.6 \text{ kcal}}{2 \text{ mol C}_3\text{H}_6\text{O}} \times 4.5 \text{ mol C}_3\text{H}_6\text{O} = -1920.6 \text{ kcal}$$

Exothermic

4. The specific heat of iron is 0.110 cal / g · °C. How much heat must be added to 5.2 g of iron at 25 °C to raise its temperature to 225 °C?

$$\text{heat} = q = m c \Delta T = (5.2 \text{ g}) \left(\frac{0.110 \text{ cal}}{\text{g} \cdot \text{C}} \right) (225 - 25 \text{ C}) = 114.4 \text{ cal}$$

Endothermic

5. Hydrazine, N₂H₄, is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate ΔH for reaction (1) given the data below.



Reverse
double quantity

$$\begin{array}{r} \text{N}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l}) \rightarrow \text{N}_2\text{H}_4(\text{l}) + \text{O}_2(\text{g}) \quad \Delta H = 622.2 \text{ kJ} \\ 2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{l}) \quad \Delta H = (2)(-285.8) \text{ kJ} \\ \hline \text{N}_2(\text{g}) + 2 \text{H}_2(\text{g}) \rightarrow \text{N}_2\text{H}_4(\text{l}) \quad \Delta H = +50.6 \text{ kJ} \end{array}$$

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S Marginson



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Table of Contents College Chemistry Practice Problems Answers

1. Understanding the eBook College Chemistry Practice Problems Answers
 - The Rise of Digital Reading College Chemistry Practice Problems Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying College Chemistry Practice Problems Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an College Chemistry Practice Problems Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from College Chemistry Practice Problems Answers
 - Personalized Recommendations
 - College Chemistry Practice Problems Answers User Reviews and Ratings
 - College Chemistry Practice Problems Answers and Bestseller Lists

5. Accessing College Chemistry Practice Problems Answers Free and Paid eBooks
 - College Chemistry Practice Problems Answers Public Domain eBooks
 - College Chemistry Practice Problems Answers eBook Subscription Services
 - College Chemistry Practice Problems Answers Budget-Friendly Options
6. Navigating College Chemistry Practice Problems Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - College Chemistry Practice Problems Answers Compatibility with Devices
 - College Chemistry Practice Problems Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of College Chemistry Practice Problems Answers
 - Highlighting and Note-Taking College Chemistry Practice Problems Answers
 - Interactive Elements College Chemistry Practice Problems Answers
8. Staying Engaged with College Chemistry Practice Problems Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers College Chemistry Practice Problems Answers
9. Balancing eBooks and Physical Books College Chemistry Practice Problems Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection College Chemistry Practice Problems Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine College Chemistry Practice Problems Answers
 - Setting Reading Goals College Chemistry Practice Problems Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of College Chemistry Practice Problems Answers
 - Fact-Checking eBook Content of College Chemistry Practice Problems Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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