

CHEMISTRY REVIEW: KEY CHAPTER 9 – CHEMICAL NAMES AND FORMULAS

- In a Stock name such as iron(II) sulfide, the roman numeral tells us (2)
 - How many atoms of Fe are in one formula unit.
 - How many sulfide ions can be attached to the iron atom.
 - The charge on each Fe ion.
 - The total positive charge of the formula unit.
- The result of changing a subscript in a correctly written chemical formula is (1)
 - Change the number of ions represented in the formula.
 - Change the charges on the other ions in the compound.
 - Change the formula so that it no longer represents the compound.
 - Have no effect on the formula.
- Derive a generalization for when an acid name will end in the suffix -ic or -ous. If ending in -ic, the anion ended in -ate. If ending in -ous, the anion ended in -ite.
- Derive a generalization for when an acid name will begin with the prefix hydro-. If prefix is hydro-, the anion ended in -ite. This is almost always a monatomic ion.
- Fill in the blanks in the table below.

Compound Name	Formula
Acetic acid	$\text{CH}_3\text{CO}_2\text{H}$
Aluminum sulfate	$\text{Al}_2(\text{SO}_4)_3$
Lead(II) carbonate	PbCO_3
Ammonium phosphate	$(\text{NH}_4)_3\text{PO}_4$
Sulfuric acid	H_2SO_4
Nitrogen pentoxide	N_2O_5
Copper(II) carbonate	CuCO_3
Sodium sulfate	Na_2SO_4
Ammonium phosphate	$(\text{NH}_4)_3\text{PO}_4$
Mercury sulfate	HgSO_4
Nitrous acid	HNO_2
Magnesium perchlorate	$\text{Mg}(\text{ClO}_4)_2$
Iron(II) nitrate	$\text{Fe}(\text{NO}_3)_2$
Iron(III) nitrate	$\text{Fe}(\text{NO}_3)_3$
Carbonic acid	H_2CO_3
Sulfur dioxide	SO_2
Sulfur pentachloride	SCl_5
Tetraphosphorus trisulfide	P_4S_3
Carbonic dioxide	CO_2

- Name the following compound and identify its molecular ions or acid.

<ol style="list-style-type: none"> $\text{Ca}(\text{OH})_2$, calcium hydroxide - II $\text{Fe}(\text{NO}_3)_2$, Iron(II) nitrate - II 	<ol style="list-style-type: none"> $\text{Fe}(\text{NO}_3)_3$, Iron(III) nitrate - III N_2O_5, Nitrogen pentoxide - III
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PT Brinkman



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High School Chemistry Michael A. Buben, 2018-05-18 The Teacher's Edition is designed to add direction to High School Chemistry which is an outline of notes. The books go hand in hand. Modified Second Edition as a student's workbook and guidance from the Teacher's Edition. **Basic Chemical Principles** Donada Peters, Edward I. Peters, 1988 **General, Organic, and Biological Chemistry** John R. Amend, Bradford P. Mundy, Melvin T. Arnold, 1990 Very Good No Highlights or Markup all pages are intact Journal of Chemical Education, 1929 **Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science**, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day. Relevant content, lively explorations, and a wealth of hands-on activities take students' understanding of science beyond the page and into the world around them. Now includes even more technology tools and activities to support differentiated instruction. **Chemistry: The Study of Matter** Henry Dorin, 1982 *NIOSH Pocket Guide to Chemical Hazards*, 1994 **The Chemical News and Journal of Physical Science**, 1913 The Druggists' Circular and Chemical Gazette, 1894 Includes Red book price list section title varies slightly issued semiannually 1897-1906 *Journal of the Society of Chemical Industry* Society of Chemical Industry (Great Britain), 1906 Includes list of members 1882-1902 and proceedings of the annual meetings and various supplements **The Wiswesser Line-formula Chemical Notation (WLN)** Elbert G. Smith, Peter A. Baker, 1975 **Chemical News and Journal of Industrial Science**, 1882 Chemical Age, 1929 **American Professional Pharmacist**, 1947 **The Chemical News and Journal of Industrial Science**, 1913 The American Perfumer, 1939 *English Mechanic and World of Science*, 1884 **Chemical News and Journal of Physical Science**, 1883 **British Chemical Abstracts**, 1926 **Merck's Report** Theodore Weicker, 1894

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