

Section 12.1

OBJECTIVE 1

Write complete balanced redox reactions that are indicated.

1. _____
 - a. _____
 - (a) balance in terms of all the atoms and products.
 - (b) indicate oxidation and reduction half-reactions.
 - (c) number of moles of each element in each compound in reaction.
 - (d) number of electrons transferred in reaction.
 - b. _____
 - (a) identify the oxidizing and reducing agents and describe the reaction.
 - (b) the name of the product of the reaction.
 - (c) the number of moles of hydrogen that will react with a certain quantity of oxygen.
 - (d) the balanced chemical equation for the reaction.
 - c. _____
 - (a) identify the oxidizing and reducing agents and describe the reaction.
 - (b) the name of the product of the reaction.
 - (c) the number of moles of each element in each compound in reaction.
 - (d) the number of electrons transferred in reaction.
 - d. _____
 - (a) identify the oxidizing and reducing agents and describe the reaction.
 - (b) the name of the product of the reaction.
 - (c) the number of moles of each element in each compound in reaction.
 - (d) the number of electrons transferred in reaction.

2. Write the balanced equation for the following redox reaction:



a. Balance the equation:



b. Write the oxidation and reduction half-reactions.

c. Identify the oxidizing and reducing agents and describe the reaction.

d. Write the balanced equation for the reaction of Mg^{2+} ions and H_2 gas.

e. Write the balanced equation for the reaction of Mg^{2+} ions and H_2 gas.

Chapter 9 Review Stoichiometry Answers Section 1

VM Jensen



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