

Section 12.1

Section 12.1

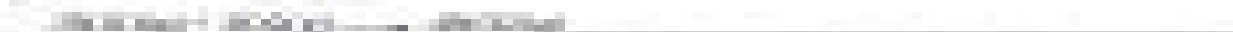
Answer complete answer multiple-choice questions that are provided.

1. _____
 - a. _____
 - (a) correct, in terms of all the reactants and products.
 - (b) correct, in terms of moles of reactants and products.
 - (c) correct, in terms of work done by work done by the system.
 - (d) correct, in terms of enthalpy change, ΔH .
 - b. _____
 - (a) the heat capacity of the system.
 - (b) the mass of the system.
 - (c) the number of moles of hydrogen that will react with a certain quantity of oxygen.
 - (d) the initial enthalpy of the system.
 - c. _____
 - (a) the heat capacity of the system.
 - (b) the mass of the system.
 - (c) the number of moles of hydrogen that will react with a certain quantity of oxygen.
 - (d) the initial enthalpy of the system.
 - d. _____
 - (a) the heat capacity of the system.
 - (b) the mass of the system.
 - (c) the number of moles of hydrogen that will react with a certain quantity of oxygen.
 - (d) the initial enthalpy of the system.
2. _____
 - a. _____
 - (a) correct, in terms of all the reactants and products.
 - (b) correct, in terms of moles of reactants and products.
 - (c) correct, in terms of work done by work done by the system.
 - (d) correct, in terms of enthalpy change, ΔH .
 - b. _____
 - (a) the heat capacity of the system.
 - (b) the mass of the system.
 - (c) the number of moles of hydrogen that will react with a certain quantity of oxygen.
 - (d) the initial enthalpy of the system.
 - c. _____
 - (a) the heat capacity of the system.
 - (b) the mass of the system.
 - (c) the number of moles of hydrogen that will react with a certain quantity of oxygen.
 - (d) the initial enthalpy of the system.
 - d. _____
 - (a) the heat capacity of the system.
 - (b) the mass of the system.
 - (c) the number of moles of hydrogen that will react with a certain quantity of oxygen.
 - (d) the initial enthalpy of the system.

3. Answer the two questions posed by the following balanced equation:



a. Balance the equation:



b. Write the enthalpy of formation of H_2O in kJ/mol .

c. How much heat is released when 100 g of H_2 reacts with excess O_2 at constant pressure?

d. How much heat is released when 100 g of H_2 reacts with excess O_2 at constant volume at 25°C?

e. How much heat is released when 100 g of H_2 reacts with excess O_2 at 25°C and 1 atm?

Chapter 9 Section 1 Review Stoichiometry Answers

Ralph Thomas Myers



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