

Chapter Test II

A. Matching

Match each item in Column I with the correct description in Column II. Write the letter of the correct description.

Column I

1. An unbalanced chemical equation for a reaction that obeys the law of conservation of matter.
2. The number of protons in an atom of an element. It is equal to the number of electrons.
3. The number of atoms of an element in a chemical formula.
4. The ratio of the actual weight of an element in a compound to its atomic weight.
5. The number of atoms of an element in a molecule of a compound.
6. The number of atoms of an element in a molecule of a compound.

Column II

- a. atomic weight
- b. limiting reagent
- c. chemical formula
- d. stoichiometry
- e. percent yield
- f. excess reagent

B. Multiple Choice

Choose the best answer for each question.

1. How many molecules of H_2O are produced when 100 g of H_2 reacts with 100 g of O_2 ?
 - a. 1000 molecules
 - b. 10000 molecules
 - c. 100000 molecules
 - d. 1000000 molecules
2. Which of the following is not a chemical change?
 - a. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
 - b. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$
 - c. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}_2$
 - d. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2$
3. The number of H_2O molecules in 100 g of H_2O is
 - a. 100
 - b. 1000
 - c. 10000
 - d. 100000

Chemistry Chapter 12 Stoichiometry Test

F Rizvi



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