

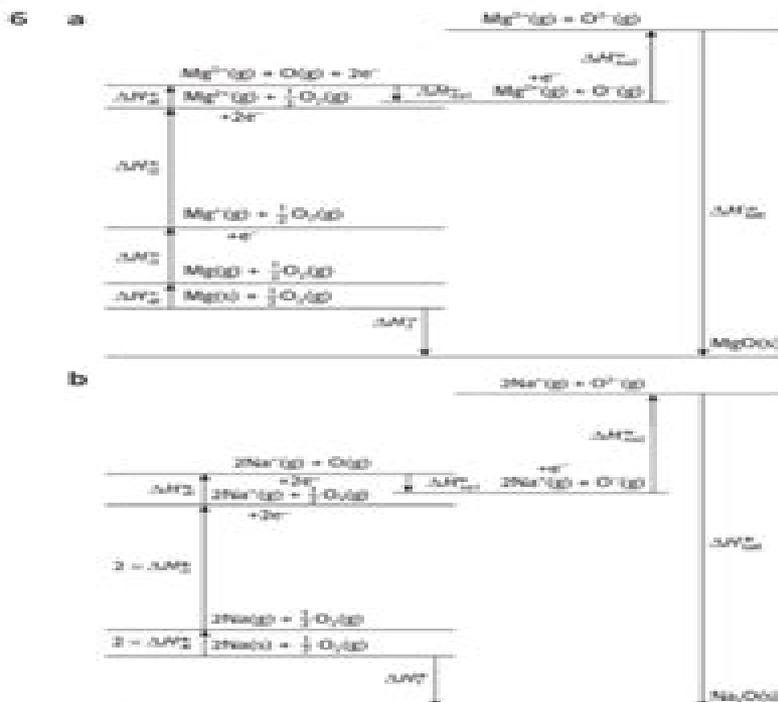
# Answers to SAQs

## Chapter 19

- 1 a standard temperature = 298 K, standard pressure =  $1 \times 10^5$  Pa / 100 kPa
- b i  $\text{Mg}^{2+}(\text{g}) + \text{O}^{2-}(\text{g}) \rightarrow \text{MgO}(\text{s})$   
 ii  $\text{K}^+(\text{g}) + \text{Br}^-(\text{g}) \rightarrow \text{KBr}(\text{s})$   
 iii  $2\text{Na}^+(\text{g}) + \text{S}^{2-}(\text{g}) \rightarrow \text{Na}_2\text{S}(\text{s})$
- 2 a The bond energy for chlorine is the enthalpy change  $\text{Cl}_2(\text{g}) \rightarrow 2\text{Cl}(\text{g})$ . The enthalpy change of atomisation is  $\frac{1}{2}\text{Cl}_2(\text{g}) \rightarrow \text{Cl}(\text{g})$ . So the enthalpy change of atomisation is  $+244/2 = +122 \text{ kJ mol}^{-1}$ .
- b i  $\frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{O}(\text{g})$   
 ii  $\text{Ba}(\text{s}) \rightarrow \text{Ba}(\text{g})$   
 iii  $\frac{1}{2}\text{Br}_2(\text{l}) \rightarrow \text{Br}(\text{g})$
- c  $0 \text{ kJ mol}^{-1}$   
 Because helium exists naturally as single gaseous atoms, no change is involved in the process  $\text{He}(\text{g}) \rightarrow \text{He}(\text{g})$ .
- 3 a There must be an input of energy to overcome the repulsive forces between the (negative) electron and the negative ion.
- b  $+440 \text{ kJ mol}^{-1}$
- c i  $\text{I}(\text{g}) + \text{e}^- \rightarrow \text{I}^-(\text{g})$   
 ii  $\text{S}^-(\text{g}) + \text{e}^- \rightarrow \text{S}^{2-}(\text{g})$
- 4 a i  $\text{Cs}(\text{g}) \rightarrow \text{Cs}^+(\text{g}) + \text{e}^-$   
 ii  $\text{Al}^{3+}(\text{g}) \rightarrow \text{Al}^{2+}(\text{g}) + \text{e}^-$   
 iii  $\text{Ca}(\text{s}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{CaO}(\text{s})$   
 iv  $\text{Fe}(\text{s}) + 1\frac{1}{2}\text{Cl}_2(\text{g}) \rightarrow \text{FeCl}_3(\text{s})$
- b  $\Delta H_{\text{form}}^{\ominus} = \Delta H_{\text{at}}^{\ominus} - [\Delta H_{\text{ion}}^{\ominus}(\text{Na}) + \Delta H_{\text{ion}}^{\ominus}(\text{Cl}) + \Delta H_{\text{at}}^{\ominus}(\frac{1}{2}\text{Cl}_2(\text{g})) + \Delta H_{\text{form}}^{\ominus}(\text{Cl})]$   
 $\Delta H_{\text{form}}^{\ominus} = (-411) - [(+107) + (+496) + (+122) + (-348)]$   
 $= +377 \text{ kJ mol}^{-1}$   
 $\Delta H_{\text{form}}^{\ominus} = (-411) - (+377) = -788 \text{ kJ mol}^{-1}$



- b i enthalpy change of atomisation of iodine  
 ii first electron affinity of nitrogen  
 iii enthalpy change of formation of strontium chloride  
 iv lattice energy of cadmium chloride



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**John T. Moore, Richard H. Langley**



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## **Table of Contents Chemistry Chapter 19 Assessment Answers**

1. Understanding the eBook Chemistry Chapter 19 Assessment Answers
  - The Rise of Digital Reading Chemistry Chapter 19 Assessment Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 19 Assessment Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Chapter 19 Assessment Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 19 Assessment Answers
  - Personalized Recommendations
  - Chemistry Chapter 19 Assessment Answers User Reviews and Ratings
  - Chemistry Chapter 19 Assessment Answers and Bestseller Lists

5. Accessing Chemistry Chapter 19 Assessment Answers Free and Paid eBooks
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  - Chemistry Chapter 19 Assessment Answers eBook Subscription Services
  - Chemistry Chapter 19 Assessment Answers Budget-Friendly Options
6. Navigating Chemistry Chapter 19 Assessment Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Chapter 19 Assessment Answers Compatibility with Devices
  - Chemistry Chapter 19 Assessment Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Chapter 19 Assessment Answers
  - Highlighting and Note-Taking Chemistry Chapter 19 Assessment Answers
  - Interactive Elements Chemistry Chapter 19 Assessment Answers
8. Staying Engaged with Chemistry Chapter 19 Assessment Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Chapter 19 Assessment Answers
9. Balancing eBooks and Physical Books Chemistry Chapter 19 Assessment Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Chapter 19 Assessment Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 19 Assessment Answers
  - Setting Reading Goals Chemistry Chapter 19 Assessment Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 19 Assessment Answers
  - Fact-Checking eBook Content of Chemistry Chapter 19 Assessment Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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