

Chapter 17: Chemical Equilibrium

Worksheet: K_c and K_p

1. For the reaction $2\text{NO}_2(g) \rightleftharpoons \text{N}_2\text{O}_4(g)$

- 1. The equilibrium constant K_c for the reaction $2\text{NO}_2(g) \rightleftharpoons \text{N}_2\text{O}_4(g)$ is 10.0 at 298 K. Calculate K_p for this reaction at 298 K.
- 2. For the reaction $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$, $K_c = 0.50$ at 400 K. Calculate K_p for this reaction at 400 K.
- 3. For the reaction $\text{H}_2(g) + \text{I}_2(g) \rightleftharpoons 2\text{HI}(g)$, $K_c = 50.0$ at 400 K. Calculate K_p for this reaction at 400 K.
- 4. For the reaction $\text{CO}_2(g) + \text{H}_2(g) \rightleftharpoons \text{CO}(g) + \text{H}_2\text{O}(g)$, $K_c = 1.60$ at 1000 K. Calculate K_p for this reaction at 1000 K.
- 5. For the reaction $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$, $K_c = 4.00$ at 1500 K. Calculate K_p for this reaction at 1500 K.
- 6. For the reaction $\text{N}_2(g) + 2\text{O}_2(g) \rightleftharpoons 2\text{NO}_2(g)$, $K_c = 1.00$ at 1500 K. Calculate K_p for this reaction at 1500 K.
- 7. For the reaction $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$, $K_c = 0.50$ at 400 K. Calculate K_p for this reaction at 400 K.
- 8. For the reaction $\text{H}_2(g) + \text{I}_2(g) \rightleftharpoons 2\text{HI}(g)$, $K_c = 50.0$ at 400 K. Calculate K_p for this reaction at 400 K.
- 9. For the reaction $\text{CO}_2(g) + \text{H}_2(g) \rightleftharpoons \text{CO}(g) + \text{H}_2\text{O}(g)$, $K_c = 1.60$ at 1000 K. Calculate K_p for this reaction at 1000 K.
- 10. For the reaction $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$, $K_c = 4.00$ at 1500 K. Calculate K_p for this reaction at 1500 K.

Answers

- 1. $K_p = 10.0 \times (RT)^{-1} = 10.0 \times (0.0821 \times 298)^{-1} = 0.406$
- 2. $K_p = K_c (RT)^{-2} = 0.50 \times (0.0821 \times 400)^{-2} = 0.00194$
- 3. $K_p = K_c (RT)^{2} = 50.0 \times (0.0821 \times 400)^2 = 1100$
- 4. $K_p = K_c (RT)^{0} = 1.60$
- 5. $K_p = K_c (RT)^{2} = 4.00 \times (0.0821 \times 1500)^2 = 390$
- 6. $K_p = K_c (RT)^{-2} = 1.00 \times (0.0821 \times 1500)^{-2} = 6.1 \times 10^{-5}$
- 7. $K_p = K_c (RT)^{-2} = 0.50 \times (0.0821 \times 400)^{-2} = 0.00194$
- 8. $K_p = K_c (RT)^{2} = 50.0 \times (0.0821 \times 400)^2 = 1100$
- 9. $K_p = K_c (RT)^{0} = 1.60$
- 10. $K_p = K_c (RT)^{2} = 4.00 \times (0.0821 \times 1500)^2 = 390$

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