

## DATA TABLE

### A. Empirical Formula of Magnesium Oxide

1. Mass of crucible plus magnesium	<u>25.748</u> g
2. Mass of crucible	<u>25.503</u> g
3. Mass of magnesium (calculate)	<u>0.245</u> g
4. Mass of crucible plus magnesium oxide	<u>25.910</u> g
5. Mass of reacted oxygen (calculate)	<u>0.967</u> g

Calculation of empirical formula (See Example 2)



Moles Mg:  $0.245 \text{ g} \times \frac{1 \text{ mol Mg}}{24.31 \text{ g Mg}} = 0.01008 \text{ mol Mg}$

Moles O:

Empirical Formula: Mg \_\_\_\_\_

# Empirical Formula Of Magnesium Oxide Report Solution

**J Spring**



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web charles s law wikipedia an animation demonstrating the relationship between volume and temperature part of a series on continuum mechanics fick s laws of diffusion laws solid mechanics fluid mechanics fluids statics dynamics archimedes principle bernoulli s principle navier stokes equations poiseuille equation pascal s law

charles law problem unacademy - Feb 26 2022

web solution  $v_1 = 2.50 \text{ l}$   $v_2 = ?$   $t_1 = 273 \text{ K}$   $t_2 = 273 + 70 = 343 \text{ K}$  we will be applying charles law as pressure is constant here  $v_1 / t_1 = v_2 / t_2$  or  $v_2 = v_1 \cdot t_2 / t_1 = 2.50 \text{ l} \cdot 343 \text{ K} / 273 \text{ K} = 3.12 \text{ l}$  at 90 c a helium sample has a volume of 500 ml determine the temperature at which the volume of the liquid will become 240 ml