

9/2/16

S. Sai Krishna

Date \_\_\_\_\_  
Page \_\_\_\_\_  
ClassmateThermochemistry

$$Q = \Delta E - W$$

$$Q = \Delta E - (-P\Delta V)$$

$$Q = \Delta E + P\Delta V$$

- At constant Volume.

$$Q = \Delta E$$

- At const. Pressure.

$$Q_p = \Delta E + P\Delta V$$

$$H = \Delta E + P\Delta V$$

$$H = \Delta E + \Delta nRT$$

$\Delta n =$  no. of moles of Prod  
- no. of moles of react.

→ For  $\Delta H$ .

- If  $\Delta H$  is -ve. reaction is exothermic.
- If  $\Delta H$  is +ve. reaction is endothermic.

Q. The heat of combustion of ethane at 17°C and const volume is  $-322.1 \text{ kJ}$ . Calculate the heat of combustion at constant pressure.



# Engineering Chemistry 2 Notes

**SA Dillow**



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### **Table of Contents Engineering Chemistry 2 Notes**

1. Understanding the eBook Engineering Chemistry 2 Notes
  - The Rise of Digital Reading Engineering Chemistry 2 Notes
  - Advantages of eBooks Over Traditional Books
2. Identifying Engineering Chemistry 2 Notes
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Engineering Chemistry 2 Notes
  - User-Friendly Interface
4. Exploring eBook Recommendations from Engineering Chemistry 2 Notes
  - Personalized Recommendations

- Engineering Chemistry 2 Notes User Reviews and Ratings
- Engineering Chemistry 2 Notes and Bestseller Lists
- 5. Accessing Engineering Chemistry 2 Notes Free and Paid eBooks
  - Engineering Chemistry 2 Notes Public Domain eBooks
  - Engineering Chemistry 2 Notes eBook Subscription Services
  - Engineering Chemistry 2 Notes Budget-Friendly Options
- 6. Navigating Engineering Chemistry 2 Notes eBook Formats
  - ePub, PDF, MOBI, and More
  - Engineering Chemistry 2 Notes Compatibility with Devices
  - Engineering Chemistry 2 Notes Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Engineering Chemistry 2 Notes
  - Highlighting and Note-Taking Engineering Chemistry 2 Notes
  - Interactive Elements Engineering Chemistry 2 Notes
- 8. Staying Engaged with Engineering Chemistry 2 Notes
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Engineering Chemistry 2 Notes
- 9. Balancing eBooks and Physical Books Engineering Chemistry 2 Notes
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Engineering Chemistry 2 Notes
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Engineering Chemistry 2 Notes
  - Setting Reading Goals Engineering Chemistry 2 Notes
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Engineering Chemistry 2 Notes
  - Fact-Checking eBook Content of Engineering Chemistry 2 Notes

- Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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