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 $2 \text{ CO}_2 \text{ (g)} + \text{H}_2\text{O (g)} \rightarrow \text{C}_2\text{H}_2 \text{ (g)} + \frac{5}{2} \text{O}_2 \text{ (g)}$ $\Delta h = 2 \text{ H}_2 \text{ (g)} + \text{C}_2\text{H}_6 \text{ (g)}$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O (g)} + \text{H}_2 \text{ (g)} + \frac{1}{2} \text{O}_2 \text{ (g)}$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ (g)} + \frac{7}{2} \text{O}_2 \text{ (g)}$ $\Delta h = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ (g)}$

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 $\text{C}_2\text{H}_4 \text{ (g)} + 3 \text{ O}_2 \text{ (g)} \rightarrow 2 \text{ CO}_2 \text{ (g)} + 2 \text{ H}_2\text{O (l)}$ $\Delta h = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6 \text{ (g)} + \frac{7}{2} \text{ O}_2 \text{ (g)} \rightarrow 2 \text{ CO}_2 \text{ (g)} + 3 \text{ H}_2\text{O (l)}$ $\Delta h = 1560 \text{ kJ/mole}$
 $\text{H}_2 \text{ (g)} + \frac{1}{2} \text{ O}_2 \text{ (g)} \rightarrow \text{H}_2\text{O (l)}$ $\Delta h = 285 \text{ kJ/mole}$

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web thus for reaction 9 2 in which 1 mole of H_2SO_4 and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = 55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions 1 $\text{NaOH}(\text{s}) + \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq})$ $\Delta H = 1$ 2 $\text{NaOH}(\text{s}) + \text{H}^+(\text{aq}) + \text{Cl}^-(\text{aq})$ $\Delta H = 2$ 3