

Chemistry - Chapter 9 Assessment Answers (Pages 281-282)

42. a) 2- b) 1+ c) 1- d) 3+

43. a) 2+ b) 2+ c) 3+ d) 1+

44. a) barium ion b) iodide ion c) silver ion d) mercury(II) ion

45. Cyanide, CN^- and hydroxide, OH^-

46. a) hydroxide ion b) lead(IV) ion c) sulfate ion d) oxide ion

47. zero; A compound is electrically neutral.

48. The symbols for the cation and anion are written and the charges are balanced with subscripts. The name of the cation is followed by the name of the anion.

49. Determine the charge of the anion then work the formula backwards to find the charge of the transition metal cation needed to give a net charge of zero for the formula unit.

50. The symbols for the cation and anion are written and the charges are balanced with subscripts. Parentheses are used around the polyatomic ion if a subscript is needed. The name of the cation is followed by the name of the anion.

51. a and b

52. When more than a single polyatomic ion is needed to balance the formula.

53. NH_4NO_3 , ammonium nitrate; $(\text{NH}_4)_2\text{CO}_3$, ammonium carbonate; NH_4CN , ammonium cyanide; $(\text{NH}_4)_3\text{PO}_4$, ammonium phosphate; $\text{Sn}(\text{NO}_3)_4$, tin(IV) nitrate; $\text{Sn}(\text{CO}_3)_2$, tin(IV) carbonate; $\text{Sn}(\text{CN})_4$, tin(IV) cyanide; $\text{Sn}_2(\text{PO}_4)_4$, tin(IV) phosphate; $\text{Fe}(\text{NO}_3)_3$, iron(III) nitrate; $\text{Fe}_2(\text{CO}_3)_3$, iron(III) carbonate; $\text{Fe}(\text{CN})_3$, iron(III) cyanide; FePO_4 , iron(III) phosphate; $\text{Mg}(\text{NO}_3)_2$, magnesium nitrate; MgCO_3 , magnesium carbonate; $\text{Mg}(\text{CN})_2$, magnesium cyanide; $\text{Mg}_3(\text{PO}_4)_2$, magnesium phosphate

54. two nonmetals

55. a) tri- b) mono- c) di- d) hexa- e) penta- f) tetra-

56. For formulas, write the correct symbols for each element with a subscript corresponding to the prefix before each element in the name. For naming, name each element in the order given. Use the subscript to determine the prefixes before each element in the name. The name ends in *-ide*.

57. a) BCl_3 b) dinitrogen pentoxide c) N_2H_4 d) carbon tetrachloride

58. a) hydrochloric acid b) H_2SO_4 c) nitric acid d) $\text{HC}_2\text{H}_3\text{O}_2$

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Table of Contents Mastering Chemistry Chapter 9 Answers

1. Understanding the eBook Mastering Chemistry Chapter 9 Answers
 - The Rise of Digital Reading Mastering Chemistry Chapter 9 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mastering Chemistry Chapter 9 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mastering Chemistry Chapter 9 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mastering Chemistry Chapter 9 Answers
 - Personalized Recommendations
 - Mastering Chemistry Chapter 9 Answers User Reviews and Ratings
 - Mastering Chemistry Chapter 9 Answers and Bestseller Lists

5. Accessing Mastering Chemistry Chapter 9 Answers Free and Paid eBooks
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 - Mastering Chemistry Chapter 9 Answers eBook Subscription Services
 - Mastering Chemistry Chapter 9 Answers Budget-Friendly Options
6. Navigating Mastering Chemistry Chapter 9 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Mastering Chemistry Chapter 9 Answers Compatibility with Devices
 - Mastering Chemistry Chapter 9 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mastering Chemistry Chapter 9 Answers
 - Highlighting and Note-Taking Mastering Chemistry Chapter 9 Answers
 - Interactive Elements Mastering Chemistry Chapter 9 Answers
8. Staying Engaged with Mastering Chemistry Chapter 9 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mastering Chemistry Chapter 9 Answers
9. Balancing eBooks and Physical Books Mastering Chemistry Chapter 9 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mastering Chemistry Chapter 9 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mastering Chemistry Chapter 9 Answers
 - Setting Reading Goals Mastering Chemistry Chapter 9 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mastering Chemistry Chapter 9 Answers
 - Fact-Checking eBook Content of Mastering Chemistry Chapter 9 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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