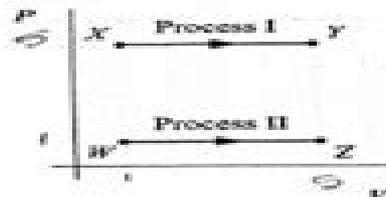




1. For each of the following parts, determine whether there exists an ideal gas process that satisfies the given conditions. If so, describe an experiment (e.g., one involving a cylinder sealed with a movable, frictionless piston) that is consistent with the given conditions. Give an example from tutorial if possible. If not, explain how you know that such a process does not exist.
  - a. There is heat transfer, but the temperature of the gas does not change (i.e.,  $Q \neq 0, \Delta T = 0$ ).  
 In the scenario via container of gas via piston, if heat is applied but the temperature doesn't change the piston would have to move up to allow more work to be done increasing the volume.
  - b. There is no heat transfer, but the temperature of the gas changes (i.e.,  $Q = 0, \Delta T \neq 0$ ).  
 A container with a gas & a piston, the piston would move down and compress the gas & increase the temperature.
  - c. There is no heat transfer, but work is done on the gas (i.e.,  $Q = 0, W \neq 0$ ).  
 The piston compressing the gas is work done on the gas.
  - d. There is heat transfer, but no work is done on the gas (i.e.,  $Q \neq 0, W = 0$ ).  
 The piston remains locked and an outside source heats the container.

2. One mole of an ideal gas is confined to a container with a movable piston. The questions below refer to the processes shown on the PV diagram at right. Process I is a change from state X to state Y at constant pressure. Process II is a change from state W to state Z at a different constant pressure.



- a. Rank states W-Z according to the temperature, from largest to smallest. If any temperatures are equal, state so explicitly. Explain.

$W(1X^2) = 1, X(1X^2) = 5, Y(5X^2) = 25, Z(5X^2) = 5,$

$Y > X = Z > W$

- b. For each process, state whether the piston moves inward, moves outward, or does not move. Explain.

$W \rightarrow Z$ : Piston moves outward (volume increases, pressure stays the same)

$X \rightarrow Y$ : Piston moves outward (volume increases, pressure remains the same)

- c. Based on your answer to part b, state whether the following quantities are positive, negative, or zero. Explain your reasoning by referring to a force and a displacement.

- i. the work done on the gas during process I ( $W_I$ )

For the volume to increase, but pressure remains the same. The process must be heated by something. Work is negative

- ii. the work done on the gas during process II ( $W_{II}$ )

This is the same thing as process 1, but at a different constant pressure, so my answer is the same as above (c-i)

- d. During process I, is the heat transfer to the gas positive, negative, or zero? Explain.

Negative, this piston would be moving outward.

# Tutorials In Introductory Physics Homework Answer Key

**Minjie Lin**



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