

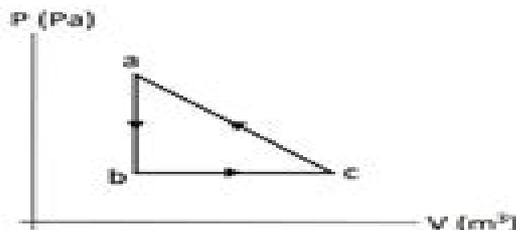
# First Law of Thermodynamics

Grade

Name \_\_\_\_\_ Date \_\_\_\_\_

Use the following P-V diagram to answer questions about this thermodynamic cycle consisting of three processes labeled  $ab$ ,  $bc$ , and  $ca$ .

$P_a = 3.00 \times 10^5 \text{ Pa}$ ,  $P_b = 1.00 \times 10^5 \text{ Pa}$ ,  $V_a = 3 \times 10^{-3} \text{ m}^3$ ,  $V_c = 9 \times 10^{-3} \text{ m}^3$ .



a) Identify the type of thermodynamic processes for each part of the cycle ( $ab$ ,  $bc$ ,  $ca$ ) as:

- |        |                |                  |                |
|--------|----------------|------------------|----------------|
| $ab$ : | i) adiabatic   | ii) isobaric     | iii) isochoric |
|        | iv) isothermal | v) none of these |                |
| $bc$ : | i) adiabatic   | ii) isobaric     | iii) isochoric |
|        | iv) isothermal | v) none of these |                |
| $ca$ : | i) adiabatic   | ii) isobaric     | iii) isochoric |
|        | iv) isothermal | v) none of these |                |

b) For the process  $ab$ , determine (i) the work  $W_{ab}$  done by the gas and (ii) the change in internal energy of the system for a heat loss of 900 J. [Hints: for (i) recall that  $W$  is the area under the curve; for (ii) use the 1<sup>st</sup> Law.]

c) For the process  $bc$ , determine (i) the work  $W_{bc}$  done by the gas and (ii) the change in internal energy of the system for a heat gain of 1250 J.

d) For the process  $ca$ , determine (i) the work  $W_{ca}$  done by the gas and (ii) the change in internal energy of the system for a heat gain of 150 J.

e) Determine the total work  $W_{tot}$  done by the gas in the cycle. Compare this to the area of the triangle. Note that both are negative. Also note that the total change in internal energy is zero, as it must be when the system returns to the same state.

# Thermodynamic Questions And Answers

**Mark Steyvers**



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