

## Question 1

1(a) Balance (top loading), spatula, stirring rod, dropper, pipette (10 cm<sup>3</sup>), pipette (25 cm<sup>3</sup>), glass (filter) funnel, burette.

1(b)

Burette readings (cm <sup>3</sup> )	Titration 1	Titration 2	Titration 3
Final volume	2.00	1.85	1.50
Initial volume	21.95	21.85	19.50
Volume of solution used	19.95*	20.00*	18.00

1(c)(i) See table above.

1(c)(ii) Average volume = (volume used in titration 1 + volume used in titration 2)/2.

Average volume =  $\{(19.95 + 20.00)/2\}$  cm<sup>3</sup>. i.e. 19.98 cm<sup>3</sup>.

Actual accurate value to be used in subsequent calculations is 19.975 cm<sup>3</sup>.

1(d) 1000 cm<sup>3</sup> K<sub>2</sub>Cr<sub>2</sub>O<sub>7(aq)</sub> contains 0.020 mol. K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub>. 19.98 cm<sup>3</sup> K<sub>2</sub>Cr<sub>2</sub>O<sub>7(aq)</sub> contains  $(19.98 \times (0.020/1000))$  mol. K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub>, i.e.  $4.00 \times 10^{-4}$  mol K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub>.

Actual accurate value to be used in subsequent calculations is  $3.995 \times 10^{-4}$  mol K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub>.

1(e) From the balanced equation, the Fe<sup>2+</sup>:Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup> stoichiometric ratio is 6:1. So 6 moles of iron ions (Fe<sup>2+</sup>) react with one mole of dichromate ions (Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup>).

1(f) Moles of iron in sample = (6 × moles of Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup> used in titration) i.e.  $(6 \times 4.00 \times 10^{-4})$  mol. =  $2.4 \times 10^{-3}$  mol. iron.

Actual accurate value to be used in subsequent calculations is  $2.397 \times 10^{-3}$  mol.

1(g) Mass of iron in salt sample = (moles of iron in sample × molar mass of iron).

Mass of iron in salt sample =  $(2.4 \times 10^{-3} \times 55.8)$  g iron i.e. 0.134 g iron.

Actual accurate value to be used in subsequent calculations is 0.1337526 g.

1(h) Percentage of iron in iron salt sample = (mass of iron in sample/mass of sample) × 100%.

Percentage of iron in iron salt sample =  $(0.134/0.500) \times 100\%$  i.e. 26.80% iron.

If accurate values were used in all calculation steps leading to this one, the final answer would be 26.75%.

1(i) Reduction of a species is the gain of an electron, or electrons, by that species.

1(j) A reducing agent is a species which causes the reduction of another species, and is itself oxidized when it reacts; it loses an electron, or electrons.

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