

## Lab 22A

### Investigating Chemical Equilibrium

**Purpose:** To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and make, via the system, visible use of LeChatelier's principle through applying LeChatelier's principle.

#### Proble:

1. Write a balanced equation for the reaction in part 1.  
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
2. Write a balanced equation for the reaction in part 2.  
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
3. Write a balanced equation for the reaction in part 3.  
$$\text{FeSCN}^{2+}(\text{aq}) + \text{H}^{+}(\text{aq}) \rightleftharpoons \text{FeSCNH}^{3+}(\text{aq})$$
4. In part 4, what is the purpose of adding potassium thiocyanate to 100 mL of 0.1 M FeCl<sub>3</sub> solution?  
To observe shifts in the equilibrium and observe the effect of adding a reactant to the reaction. The purpose of adding FeSCN<sup>2+</sup> is to observe the effect of adding a reactant to the reaction and observe the effect of adding a reactant to the reaction.
5. What is the purpose of adding distilled water to part 4?  
To observe the effect of adding a reactant to the reaction and observe the effect of adding a reactant to the reaction.
6. What is the purpose of adding 10 mL of 0.1 M FeCl<sub>3</sub> to part 4?  
To observe the effect of adding a reactant to the reaction and observe the effect of adding a reactant to the reaction.
7. Write the equation for the reaction between Fe<sup>3+</sup> and OH<sup>-</sup> ions.  
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})_3(\text{s})$$
8. Write the equation for the reaction between Fe<sup>3+</sup> and OH<sup>-</sup> ions.  
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})_3(\text{s})$$
9. What is the purpose of adding 10 mL of 0.1 M FeCl<sub>3</sub> to part 4?  
To observe the effect of adding a reactant to the reaction and observe the effect of adding a reactant to the reaction.

# Investigating Chemical Equilibrium Lab 12a Answer

**Anthony S. Fauci**



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