

# CBGS SCHEME

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15EC42



## Fourth Semester B.E. Degree Examination, July/August 2021 Microprocessor

Time: 3 Hrs.

Max. Marks: 80

Note: Answer any FIVE full questions.

1. a. With a relevant diagram, explain the register organization of 8086. (08 Marks)  
b. List any three the advantages of memory segmentation. (03 Marks)  
c. The contents of different registers are shown below. Form the effective addresses for the instructions given below. Offset (displacement) = 5000H.  
[AX] = 1000H, [BX] = 2000H, [SI] = 3000H, [DI] = 4000H, [BP] = 5000H, [SP] = 6000H, [CS] = 0000H, [DS] = 1000H, [SS] = 2000H, [IP] = 7000H.  
i) MOV AX, [5000H]  
ii) MOV AX, [BX]  
iii) MOV AX, 5000H[BX]  
iv) MOV AX, [BX][SI]  
v) MOV AX, 5000H [BX] [SI]. (05 Marks)
2. a. Explain PUSH AX and POP AX instructions with steps involved. (05 Marks)  
b. Write an 8086 ALP to multiply two 8 bit signed numbers. Give example for different cases. (06 Marks)  
c. If the machine code for OPCODE MOV is 100010, find the hex codes for,  
i) MOV AX, BX  
ii) MOV AX, [BX] (05 Marks)
3. a. Explain different string handling instructions. (05 Marks)  
b. Write an 8086 ALP for the addition of two  $3 \times 3$  matrices shown below. The matrices are stored in the form of lists (row wise). Store the result of addition in the third list.  
$$\text{MAT1} = \begin{bmatrix} 01\text{H} & 02\text{H} & 03\text{H} \\ 04\text{H} & 05\text{H} & 06\text{H} \\ 07\text{H} & 08\text{H} & 09\text{H} \end{bmatrix} \quad \text{MAT2} = \begin{bmatrix} 11\text{H} & 22\text{H} & 33\text{H} \\ 44\text{H} & 55\text{H} & 66\text{H} \\ 77\text{H} & 88\text{H} & 99\text{H} \end{bmatrix}$$
 (05 Marks)
- c. Explain the following instructions with example  
i) AAA ii) SAR iii) NEG. (06 Marks)
4. a. Explain the following assembler directives with example.  
i) ORG ii) PTR iii) PROC. (06 Marks)  
b. Write an 8086 ALP to convert an 8 bit binary number into equivalent ECD code. (05 Marks)  
c. If [CL] = 56, find the contents of register BL, after execution of following set of instructions  
MOV BL, 1  
MOV AL, 0  
UP : CMP CL, 0  
JZ END  
SUB CL, BL  
INC AL  
ADD BL, 02  
JMP UP  
END : MOV BL, AL (05 Marks)

Important Note : 1. On completing your answers, compulsorily draw diagonal cross lines on the remaining blank pages.  
2. Any revealing of identification, appeal to evaluator and/or equisation written eg. 42-V-30, will be treated as malpractice.

# Microprocessor Question Paper Vtu

**Annelies Wilder-Smith**



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this video accompanies the chm 252 lab nitration of bromobenzene in this video students will get an introduction to the lab as well as the set up for conducting the reaction students should

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for the nitration 4.01 ml of nitric acid 4.00 ml of sulfuric acid and 4.56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4.56 ml of bromobenzene was used the amount of product predicted to be recovered was 8.7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction especially nitration of methyl benzoate to give methyl m nitrobenzoate thus determine the weight melting point and percentage yield of the pure methyl m nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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1 go to the restroom and wash the affected area with lots of soap and water the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area 3 the summary to this warning is not to touch any part of your body while performing this experiment if you must leave the lab for any reason 1

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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nitration of bromobenzene lab report improved essays 731 words 3 pages open document essay sample check writing quality show more this experiment was performed in order to demonstrate the principles of nitration through electrophilic aromatic substitution it also examined the effects of bromine on an aromatic ring ortho para vs meso directing

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then actual is your actual findings yield actual theoretical  $\times 100$  what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area theory and explanation of the nitration of bromobenzene if the temperature of the reaction mixture exceeds 60 o

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1 bromo 4 nitrobenzene

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study with quizlet and memorize flashcards containing terms like reacting sulfuric and nitric acid together produces the nitronium ion alone nitric acid is not a strong enough electrophile to react with bromobenzene sulfuric acid is added to the nitric acid the reverse rxn can overheat and boil easily includes a claisen adaptor and

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture hno 3 h 2 so 4 in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accommodate their electrons demonstration of creating the p arenium cation by to nitration of bromobenzene

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nitrobromobenzene recrystallization was used to purify 4 nitrobromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion  $\text{NO}_2^+$  which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4.0ml of nitric acid and 4.0ml of sulfuric acid and will then react with 4.5ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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experimental spectra reaction a fluorobenzene stock fluorobenzene 1 h 1 h 19 f 19 f 13 c nmr fid for reference only sample 1 h 1 h 19 f 19 f 13 c nmr eas nitration of fluorobenzene crude product not available for submission for credit

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chem 2261 lab 5 nitration of bromobenzene introduction an electrophilic aromatic substitution reaction is an important process that allows for intermediate chemical structures to be used as precursors to pharmaceuticals and other industrially produced chemicals by adding different functional groups into an aromatic ring

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lab report 10 nitration of bromobenzene raekwon filmore cm 244 section 40 march 27 2018 introduction for this experiment

nitration of bromobenzene was the focus of the lab the benzene is an aromatic compound and when it reacts with wither a mixture of sulfuric acid or nitric acid creates what is known as a nitro group

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web jul 18 2022 because an equilibrium state is achieved when the forward reaction rate equals the reverse reaction rate under a given set of conditions there must be a relationship between the composition of the system at equilibrium and the kinetics of a reaction represented by rate constants

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web reaction rate the number of particles that react in a given time to form products le châtelier s principle if a stress is applied to a system in dynamic equilibrium the system changes to relieve the stress chemical equilibrium the forward and reverse reactions take place at the same rate activation energy

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web what can be calculated if the concentration of a reaction s products and reactants are known as well as the coefficients of each in the balanced equation the equilibrium constant what is the equation for the equilibrium

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web in order to physically measure the rate of consumption of the bisulfite ion we will look for a color change in the reaction vessel the solution will change from clear to dark blue if you have the following equilibrium reaction  $co(g) + 2h_2(g) \rightleftharpoons ch_3oh(g)$

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equilibrium of a reversible reaction can be altered by changing the

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