

## AP Chemistry Practice Test Chapter 8 and 9

### Multiple Choice

Identify the choice that best completes the statement or answers the question.

- \_\_\_\_\_ 1. Which pair of atoms should form the most polar bond?  
a. F and B  
b. C and O  
c. F and O  
d. N and F  
e. B and N
- \_\_\_\_\_ 2. Which pair of ions should form the ionic lattice with the highest energy?  
a.  $\text{Na}^{1+}$  and  $\text{Br}^{1-}$   
b.  $\text{Li}^{1+}$  and  $\text{F}^{1-}$   
c.  $\text{Cs}^{1+}$  and  $\text{F}^{1-}$   
d.  $\text{Li}^{1+}$  and  $\text{O}^{2-}$   
e.  $\text{K}^{1+}$  and  $\text{F}^{1-}$
- \_\_\_\_\_ 3. The electrostatic force of attraction is greatest in which compound?  
a.  $\text{BaO}$   
b.  $\text{MgO}$   
c.  $\text{CaS}$   
d.  $\text{MgS}$   
e.  $\text{CaO}$
- \_\_\_\_\_ 4. Which molecule has the weakest bond?  
a.  $\text{CO}$   
b.  $\text{O}_2$   
c.  $\text{NO}$   
d.  $\text{N}_2$   
e.  $\text{Cl}_2$
- \_\_\_\_\_ 5. For which species are octet resonance structures necessary and sufficient to describe the bonding satisfactorily?  
a.  $\text{BCl}_3$   
b.  $\text{SO}_2$   
c.  $\text{CO}_2$   
d.  $\text{BeF}_2$   
e.  $\text{SO}_4^{2-}$
- \_\_\_\_\_ 6. How are the bonding pairs arranged in the best Lewis structure for ozone?  
a.  $\text{O}-\text{O}-\text{O}$   
b.  $\text{O}=\text{O}-\text{O}$   
c.  $\text{O}=\text{O}=\text{O}$   
d.  $\text{O}=\text{O}=\text{O}$   
e.  $\text{O}=\text{O}=\text{O}$
- \_\_\_\_\_ 7. Which species has the shortest bond length?  
a.  $\text{CN}^{1-}$   
b.  $\text{O}_2$   
c.  $\text{SO}_2$   
d.  $\text{SO}_3$   
e.  $\text{CO}_2$
- \_\_\_\_\_ 8. Which species requires the least amount of energy to remove an electron from the outermost energy level?  
a.  $\text{Na}^{1+}$   
b. He  
c.  $\text{F}^{1-}$   
d.  $\text{O}^{2-}$   
e.  $\text{Mg}^{2+}$
- \_\_\_\_\_ 9. Which species has a valid non-octet Lewis structure?  
a.  $\text{GeCl}_4$   
b.  $\text{SiF}_4$   
c.  $\text{CO}_3^{2-}$   
d.  $\text{SeCl}_4$   
e.  $\text{CO}_3^{2-}$

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**IM Harris**



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## **Table of Contents Ap Chem Chapter 1 Practice Test**

1. Understanding the eBook Ap Chem Chapter 1 Practice Test
  - The Rise of Digital Reading Ap Chem Chapter 1 Practice Test
  - Advantages of eBooks Over Traditional Books
2. Identifying Ap Chem Chapter 1 Practice Test
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ap Chem Chapter 1 Practice Test
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ap Chem Chapter 1 Practice Test
  - Personalized Recommendations
  - Ap Chem Chapter 1 Practice Test User Reviews and Ratings
  - Ap Chem Chapter 1 Practice Test and Bestseller Lists
5. Accessing Ap Chem Chapter 1 Practice Test Free and Paid eBooks
  - Ap Chem Chapter 1 Practice Test Public Domain eBooks
  - Ap Chem Chapter 1 Practice Test eBook Subscription Services
  - Ap Chem Chapter 1 Practice Test Budget-Friendly Options

6. Navigating Ap Chem Chapter 1 Practice Test eBook Formats
  - ePub, PDF, MOBI, and More
  - Ap Chem Chapter 1 Practice Test Compatibility with Devices
  - Ap Chem Chapter 1 Practice Test Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ap Chem Chapter 1 Practice Test
  - Highlighting and Note-Taking Ap Chem Chapter 1 Practice Test
  - Interactive Elements Ap Chem Chapter 1 Practice Test
8. Staying Engaged with Ap Chem Chapter 1 Practice Test
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ap Chem Chapter 1 Practice Test
9. Balancing eBooks and Physical Books Ap Chem Chapter 1 Practice Test
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ap Chem Chapter 1 Practice Test
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ap Chem Chapter 1 Practice Test
  - Setting Reading Goals Ap Chem Chapter 1 Practice Test
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ap Chem Chapter 1 Practice Test
  - Fact-Checking eBook Content of Ap Chem Chapter 1 Practice Test
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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