Stoichiometry Worksheet # 1

1) Glucose is used as a source of energy by the human body. The overall reaction in the body is

$$C_6H_{12}O_6 + 6O_2 \rightarrow 6CO_2 + 6H_2O$$

Calculate the number of moles of oxygen needed to oxidize 12.5 g of $C_6H_{12}O_6$.

2) Ammonia is synthesized from hydrogen and nitrogen according to the following equation.

$$N_2 + 3H_2 \rightarrow 2NH_3$$

If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia (NH_3) can be produced?

3) Assume that in the decomposition of potassium chlorate, $KCLO_3$,

$$KCLO_3 \rightarrow KCLO + O_2$$

80.5 moles of O_2 form. How many grams of potassium chlorite, the other product would be formed?

4) In a single displacement reaction,

$$2Al + 2HCl \rightarrow 2AlCl + H_2$$

9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

5) The compound "Cisplatin" $PtCl_2(NH_3)_2$, has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction

$$K_2PtCl_4 + 2NH_3 \dashrightarrow 2KCl + PtCl_2(NH_3)_2$$

A. How many moles of "cisplatin" can be produced from 2.50 g K_2PtCl_4?

B. How many moles of NH_3 would be needed if 3.8 moles of K_2PtCl_4 were present?

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