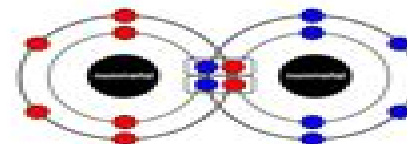


DRAWING COVALENT BONDS

Introduction:

A **covalent bond** forms between **nonmetals**. Covalent bonds involve **sharing pairs of electrons** between two atoms so that each atom gains a full outer valence shell. Sharing 1 pair of electrons creates a **single covalent bond**, sharing 2 pairs of electrons creates a **double covalent bond**, and sharing 3 pairs of electrons creates a **triple covalent bond**.



How to Draw Simple Covalent Bond Diagrams	
1.	Draw the electron dot diagrams for each element. <i>HINT: This will help you see the number of unpaired valence electrons for each element.</i>
2.	Count the total number of valence electrons between all atoms. <i>HINT: The total number of valence electrons should be 8 in most cases, except 2 hydrogens. If there are not 8, add more atoms of the element with the least number of valence electrons to start.</i>
3.	Determine which element is the central atom in the molecule, if needed. <i>HINT: The central atom is typically the atom with the largest valence number or the atom that there is only one of in the formula. If there are only 2 atoms in the formula, there is no central atom!</i>
4.	Pair up single electrons from different atoms to form covalent bonds and circle each pair. <i>HINT: Keep doing this until each single electron from all atoms have been paired up. You may need to use double or triple bonds to make this work. This step takes trial and error. If one arrangement is not working, try adding another element or trying a new structure. Practice makes perfect!</i>

Activity: Drawing Covalent Bond Diagrams:

- For each covalent compound, you have been given the element names and the chemical formula.
- Follow the steps listed above to create a covalent bonding diagram for the compound.
- Remember, this takes trial and error. The first 2 have been done for you as examples.

#	Nonmetal	Nonmetal	Formula	Bonding Diagram
A.	hydrogen H • <i>(must share 1)</i>	chlorine • Cl • <i>(must share 1)</i>	HCl <i>(Formula shows 1 H and 1 Cl)</i>	
B.	hydrogen H • <i>(must share 1)</i>	oxygen • O • <i>(must share 2)</i>	H ₂ O <i>(Formula shows 2 H and 1 O)</i>	

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M Walker



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Table of Contents Chapter 14 Covalent Bonds Review Sheet Answers

1. Understanding the eBook Chapter 14 Covalent Bonds Review Sheet Answers
 - The Rise of Digital Reading Chapter 14 Covalent Bonds Review Sheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 14 Covalent Bonds Review Sheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 14 Covalent Bonds Review Sheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 14 Covalent Bonds Review Sheet Answers
 - Personalized Recommendations
 - Chapter 14 Covalent Bonds Review Sheet Answers User Reviews and Ratings
 - Chapter 14 Covalent Bonds Review Sheet Answers and Bestseller Lists

5. Accessing Chapter 14 Covalent Bonds Review Sheet Answers Free and Paid eBooks
 - Chapter 14 Covalent Bonds Review Sheet Answers Public Domain eBooks
 - Chapter 14 Covalent Bonds Review Sheet Answers eBook Subscription Services
 - Chapter 14 Covalent Bonds Review Sheet Answers Budget-Friendly Options
6. Navigating Chapter 14 Covalent Bonds Review Sheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 14 Covalent Bonds Review Sheet Answers Compatibility with Devices
 - Chapter 14 Covalent Bonds Review Sheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 14 Covalent Bonds Review Sheet Answers
 - Highlighting and Note-Taking Chapter 14 Covalent Bonds Review Sheet Answers
 - Interactive Elements Chapter 14 Covalent Bonds Review Sheet Answers
8. Staying Engaged with Chapter 14 Covalent Bonds Review Sheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 14 Covalent Bonds Review Sheet Answers
9. Balancing eBooks and Physical Books Chapter 14 Covalent Bonds Review Sheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 14 Covalent Bonds Review Sheet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 14 Covalent Bonds Review Sheet Answers
 - Setting Reading Goals Chapter 14 Covalent Bonds Review Sheet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 14 Covalent Bonds Review Sheet Answers
 - Fact-Checking eBook Content of Chapter 14 Covalent Bonds Review Sheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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