

# Reaction Energy

## Section 1

Answer **ALL** of the following questions in the space provided.

- The following data are provided from the table of  $\Delta H_f^\circ$  values.
- The following are the combustion of hexane and benzene respectively. The following thermochemical equations:



- By using bond enthalpies or by referring to standard liquid  $\Delta H_f^\circ$  data, calculate:
- the heat of the combustion of hexane and benzene.

Answer **ALL** of the following questions in the box to the left of the question in the space provided.

- 10.000 g of water at 20.0°C absorbs a certain amount of energy when it is heated to 100.0°C.

- Calculate the specific heat of water (e.g.  $\text{J}^\circ\text{C}^{-1}\text{g}^{-1}$ ) from calorimetric data for the combustion of hexane and benzene. (The  $\Delta H_f^\circ$  values of substances from 20.0°C to 100.0°C are given in the table below.)

- The combustion of hexane and benzene are given by the following thermochemical equations:

- Calculate the heat of combustion of hexane and benzene.

# Chapter 16 Review Reaction Energy Section 1 Short Answer

**BM King**



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Chemistry Martin Stuart Silberberg,1996 This general chemistry text centres on the theme that observable change in chemical systems is the result of molecular change The aims of this edition are to enable students to perceive matter and change at the molecular level and to help build student confidence in their ability to solve chemical problems as they discover the relevance of chemistry to their lives      **Farmers' Review** ,1896      *Study Guide for Chemical Principles* Thomas Elliott Taylor,1979      Focus on physical science Charles H. Heimler,1989      **British Chemical Abstracts** ,1951

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