

Chapter 16 Solubility and Complex Ion Equilibria

Section 16.1: Solubility Equilibria and the Solubility Product

◆ 1. What is Solubility Equilibrium?

- When an **ionic solid** (like AgCl) is put in water, some dissolves into ions.
- At the same time, ions in the water can recombine and form the solid again.
- Eventually, it reaches **equilibrium**: the rate of dissolving = rate of crystallizing.

Example: $AgCl_{(s)} \rightleftharpoons Ag^+_{(aq)} + Cl^-_{(aq)}$

◆ 2. Solubility Product (K_{sp})

- **K_{sp}** is the equilibrium constant for this dissolving.
- For AgCl: $K_{sp} = [Ag^+][Cl^-]$

👉 Notice: The solid (AgCl) is left out because its concentration is constant.

◆ 3. Solubility vs. K_{sp}

- **Solubility** = how many moles per liter dissolve.
- **K_{sp}** = a number from the ion concentrations at equilibrium.

⚠ Important: A larger K_{sp} doesn't always mean higher solubility because the ion ratios matter.

Example:

- AgCl: $K_{sp} = 1.6 \times 10^{-10}$, solubility $\approx 1.3 \times 10^{-5}$.
 - Ag₃PO₄: $K_{sp} = 1.8 \times 10^{-18}$, solubility $\approx 1.6 \times 10^{-5}$.
Even though Ag₃PO₄ has a smaller K_{sp}, its solubility is close because of the 3:1 ratio of ions.
-

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