



This is the expression for the equilibrium constant for the reaction shown above.

A $K_{\text{eq}} = [\text{CO}_2][\text{H}_2]$

B $K_{\text{eq}} = \frac{[\text{CO}_2]^2[\text{H}_2]}{[\text{CO}][\text{H}_2\text{O}]^2}$

C $K_{\text{eq}} = \frac{[\text{CO}_2][\text{H}_2]}{[\text{CO}][\text{H}_2\text{O}]}$

D $K_{\text{eq}} = \frac{[\text{CO}][\text{H}_2\text{O}]}{[\text{CO}_2][\text{H}_2]}$

2. The hydrogen sulfide produced as a byproduct of petroleum refinement can be used to produce elemental sulfur:



What is the equilibrium constant expression for this reaction?

A $K_{\text{eq}} = \frac{[\text{H}_2\text{O}]^2}{[\text{H}_2\text{S}]^2[\text{SO}_2]}$

B $K_{\text{eq}} = \frac{[\text{S}]^3[\text{H}_2\text{O}]^2}{[\text{H}_2\text{S}]^2[\text{SO}_2]}$

C $K_{\text{eq}} = \frac{[\text{H}_2\text{S}]^2[\text{SO}_2]}{[\text{H}_2\text{S}]^2}$

D $K_{\text{eq}} = \frac{[\text{H}_2\text{O}]}{[\text{H}_2\text{S}][\text{SO}_2]}$

3. What can be determined if the ion product (Q_{sp}) is compared to the solubility product constant (K_{sp})?

A the law of mass action

B whether a substance will precipitate

C equilibrium

D the volume of the solution

4. The value of any equilibrium constant is correct only at

A a specific volume.

B a specific pressure.

C a specific concentration.

D a specific temperature.

5. A system reaches chemical equilibrium when

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AW Chickering



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