

## CHAPTER

## 06

## THERMODYNAMICS

- + Universe = system + surroundings
- +  $W = F \times dx$  where  $W$  = work done,  $F$  = force,  $dx$  = distance moved
- +  $W = P_{ext} \Delta V$  ( $\Delta V = V_2 - V_1$ )
- +  $dW = P dV$
- +  $W = \int_{V_1}^{V_2} P_{ext} dV$
- + Work done by a system = -ve
- + Work done on the system = +ve
- +  $C = \frac{q}{dT}$ ;  $q$  = Heat absorbed,  $dT$  = rise in temperature
- +  $Q = \Delta E + W$  ..... I law of thermodynamics
- +  $q = dE + W$  ..... I law of thermodynamics
- + Heat absorbed by a system = +ve
- + Heat given out by a system = -ve
- +  $H = E + PV$
- +  $\Delta H = \Delta E + P \Delta V$
- +  $C_v = \left( \frac{\partial E}{\partial T} \right)_v$ ;  $C_v$  = Heat capacity at constant volume  
 $\partial E$  = Change in internal energy,  $\partial T$  = change in temperature at constant volume in Kelvin
- +  $C_p = \left( \frac{\partial H}{\partial T} \right)_p$
- +  $C_p - C_v = R$ ,  $R$  = Gas constant,  $C_p$  = Heat capacity at constant pressure
- +  $\Delta H$  = Change in enthalpy,  $\partial T$  = Change in temperature at constant pressure
- +  $\frac{C_p}{C_v} = \gamma$
- + Molar heat capacity = specific heat x molecular weight
- + Heat liberating reactions are exothermic reactions,  $\Delta H = -ve$
- + Heat absorbing reactions are endothermic reactions,  $\Delta H = +ve$
- + Heat of combustion =  $Z \times \theta \times \frac{M}{m}$  (cals),  
 $M$  = molecular weight,  $m$  = weight of substance  
 $\theta$  = rise in temperature  
 $Z$  = heat capacity
- + Heat liberated = (W + volume of reaction mixture) x rise in temperature
- +  $W = \left[ \frac{m_2 (t_2 - t_3)}{t_3 - t_1} - m_1 \right]$ ,  $m_1$  = mass of water at lower temperature  
 $m_2$  = mass of water at higher temperature  
 $t_1^\circ C$  = Lower temperature,  $t_2^\circ C$  = higher temperature  
 $t_3^\circ C$  = resultant temperature,  $W$  = water equivalent of calorimeter + stirrer

# Chapter 1 Solutions Thermodynamics

**Garrison Sposito**



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