

7B-3 : Entropy Change of an Isobaric Process

6 pts

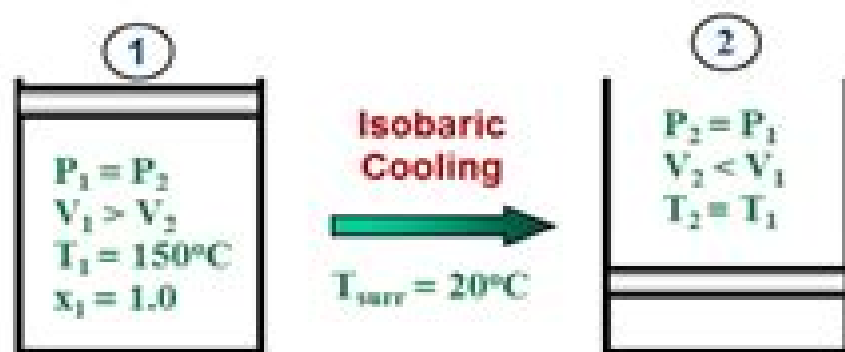
A piston-and-cylinder device with a free-floating piston contains 2.6 kg of saturated water vapor at 150°C. The water loses heat to the surroundings until the cylinder contains saturated liquid water.

The surroundings are at 20°C. Calculate... a.) ΔS_{water} , b.) $\Delta S_{\text{surroundings}}$, c.) $\Delta S_{\text{universe}}$

Read : Calculating ΔS for the water in the cylinder is straightforward. The key to calculating ΔS_{surr} is the fact that the surroundings behave as a thermal reservoir. The temperature of the surroundings does not change and there are no irreversibilities in the surroundings that are associated with the process. The key to calculating ΔS_{univ} is the fact that the universe is made up of the combination of the system and the surroundings. Consequently, $\Delta S_{\text{univ}} = \Delta S_{\text{sys}} + \Delta S_{\text{surr}}$.

Given:	m	2.60	kg	Find:	ΔS_{sys}	???	kJ/K
	T_1	150	°C		ΔS_{surr}	???	kJ/K
	x_1	1	kg vap/kg		ΔS_{univ}	???	kJ/K
	T_{surr}	20	°C				
	$P_1 = P_2$						
	x_2	0	kg vap/kg				

Diagram:



- Assumptions:**
1. Changes in kinetic and potential energies are negligible.
 2. Boundary work is the only form of work that crosses the system boundary.
 3. The surroundings behave as a thermal reservoir.

Entropy Problems And Solutions

Rachel Sandford



Entropy Problems And Solutions:

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