

STOICHIOMETRY

QUESTIONS

1. Will an increase in temperature always increase the solubility of the solute in the solvent? Why or why not?

No, because in the case of solid solutes, the effect of the temperature towards solubility depends on what kind of reaction, either exothermic or endothermic. In endothermic reaction as the temperature increases, the solubility increases as well but in exothermic reaction, if the temperature increases, the solubility decreases. The same manner with gas particles, increased temperature, makes higher solubility.

2. Why is glycerine soluble in water but not in diethyl ether?

It is because water is a polar solvent making glycerine (a polar solute) to become soluble, however diethyl ether is a non-polar solvent making glycerine insoluble.

3. How will you construct a procedure in preparation of 3 M NaCl (100mL solution)?

First, compute for the needed grams/mol of NaCl to put in the solvent. Since we need to prepare a 3 M NaCl in 100ml solution, we compute it by $0.100\text{L}(3\text{mol})/1\text{L}=0.3 \text{ mol NaCl}$ then multiplying 0.3 mol NaCl to 58.43 g/mol NaCl (molar mass of NaCl compound) resulting to 17.529 g/NaCl ,which will be the amount in weight that will be put in the solvent. Second, measure as close to 17.529 g NaCl in weighing paper using a balance.

Then transfer it to 100 ml volumetric flask. Third, put the water solvent in the in the bottom of the flask in approximately half full then shake to dissolve the NaCl solute. Lastly, carefully put more water until reaching the rest of the bulb area up to the neck of the flask then add solvent drop-wise monitoring the location of the meniscus, then labeled the flask for laboratory protocols.

● CONCLUSION

Summarized the basic properties and preparation of solutions. Solutions are known to be the homogeneous mixture where in substances and its components are uniformly distributed. Solvents are components that are greater amounts and the solute which is the lesser amount that is miscible or soluble. Solubility is

Experiment 6 Stoichiometry Lab Report Conclusion

D Kirk



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