

Practicing the Mole - - Even Problems Answer Key

Calculate the mass in grams of each of the following:

2. 8.00 moles of aluminum

$$\frac{8.00 \text{ mol} \mid 26.98 \text{ g}}{1 \text{ mol}} = 216 \text{ g Al}$$

6. 7.00 moles of iodine (I₂)

$$\frac{7.00 \text{ mol} \mid 253.8 \text{ g}}{1 \text{ mol}} = 1,780 \text{ g I}_2$$

4. 2.00 x 10² moles of chlorine (Cl₂)

$$\frac{200. \text{ mol} \mid 70.90 \text{ g}}{1 \text{ mol}} = 14,200 \text{ g Cl}_2$$

8. 9.20 moles of iron

$$\frac{9.20 \text{ mol} \mid 55.85 \text{ g}}{1 \text{ mol}} = 514 \text{ g Fe}$$

Calculate the number of moles of atoms in each of the following:

10. 281 g of iron

$$\frac{281 \text{ g} \mid 1 \text{ mol}}{55.85 \text{ g}} = 5.03 \text{ mol Fe}$$

14. 93.0 g of phosphorus

$$\frac{93.0 \text{ g} \mid 1 \text{ mol}}{30.97 \text{ g}} = 3.00 \text{ mol P}$$

12. 72.0 g of magnesium

$$\frac{72.0 \text{ g} \mid 1 \text{ mol}}{24.31 \text{ g}} = 2.96 \text{ mol Mg}$$

16. 1035 g of lead

$$\frac{1035 \text{ g} \mid 1 \text{ mol}}{207.2 \text{ g}} = 4.995 \text{ mol Pb}$$

Calculate the number of atoms in each of the following:

18. 5.40 moles of iron

$$\frac{5.40 \text{ mol} \mid 6.022 \times 10^{23} \text{ atoms}}{1 \text{ mole}}$$

$$= 3.25 \times 10^{24} \text{ atoms Fe}$$

20. 10.0 moles of titanium

$$\frac{10.0 \text{ mol} \mid 6.022 \times 10^{23} \text{ atoms}}{1 \text{ mole}}$$

$$= 6.02 \times 10^{24} \text{ atoms Ti}$$

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G Thomas



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Chemistry McGraw-Hill Staff, 2001-03 **Fundamentals of Physics, , Problem Supplement No. 1** Jearl Walker, 2001
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web the real power absorbed by the load is given by $p = 120 \cdot 6 \cdot 63 \cdot \cos 56.4^\circ = 440 \text{ W}$ which can be checked by $i_r^2 \cdot 6 \cdot 63 = 10 \cdot 440 \text{ W}$ the reactive power absorbed by the load is $q = 120 \cdot 6 \cdot 63 \cdot \sin 36.4^\circ = 663 \text{ var}$ c peak magnetic energy $w_{li} = \frac{1}{2} \cdot 0.04 \cdot 6 \cdot 63 \cdot 1.76^2 = 377.1 \text{ J}$ $q_w = \omega$

$377.1 \text{ J} = 663 \text{ var}$ is satisfied $2 \cdot 17 \text{ A} \cdot \text{svi} = 2 \cdot 2 \cdot \omega$

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