



## Chapter 12

### *Chemical Kinetics*

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### Chemical Kinetics

Thermodynamics – does a reaction take place?

Kinetics – how fast does a reaction proceed?

**Reaction rate** is the change in the concentration of a reactant or a product with time (*M/s*).



$$\text{rate} = -\frac{\Delta[A]}{\Delta t}$$

$\Delta[A]$  = change in concentration of A over time period  $\Delta t$

$$\text{rate} = \frac{\Delta[B]}{\Delta t}$$

$\Delta[B]$  = change in concentration of B over

# Chapter 12 Chemical Kinetics Answer Key

**JL Elias**



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