

**14 • The Chemistry of Acids & Bases****STUDY QUESTIONS & PROBLEMS - KEY**

1. For the following aqueous equilibria, designate the Bronsted-Lowry conjugate acid-base pairs and establish the weaker side:



2. Write the name and formula for the conjugate bases of the following:



3. Complete the Bronsted-Lowry equilibria, label the components acid or base, and pair up the conjugate acid-base pairs:



4. Is the monohydrogen phosphate ion  $\text{HPO}_4^{2-}$  amphiprotic? -- YES

If so, write the formulas of its conjugate acid and its conjugate base. CA =  $\text{H}_2\text{PO}_4^-$  and CB =  $\text{PO}_4^{3-}$

# Chapter 17 Study Acids Bases

**Katrin Zwirgmaier**



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