

Honors Chemistry
Extra Stoichiometry Problems

1. Silver nitrate reacts with barium chloride to form silver chloride and barium nitrate.

a. Write and balance the chemical equation.



b. If 39.02 grams of barium chloride are reacted in an excess of silver nitrate, how many representative particles (and what type) of silver chloride are produced?

$$\frac{39.02 \text{ g BaCl}_2}{1} \times \frac{1 \text{ mole BaCl}_2}{208.233 \text{ g BaCl}_2} \times \frac{2 \text{ moles AgCl}}{1 \text{ mole BaCl}_2} \times \frac{6.02 \times 10^{23} \text{ f. u. AgCl}}{1 \text{ mole AgCl}} = 2.256 \times 10^{23} \text{ f. u. AgCl}$$

c. If 410.8 grams of barium nitrate are produced how many grams of silver nitrate were reacted?

$$\frac{410.8 \text{ g Ba}(\text{NO}_3)_2}{1} \times \frac{1 \text{ mole Ba}(\text{NO}_3)_2}{261.335 \text{ g Ba}(\text{NO}_3)_2} \times \frac{2 \text{ moles AgNO}_3}{1 \text{ mole Ba}(\text{NO}_3)_2} \times \frac{169.872 \text{ g AgNO}_3}{1 \text{ mole AgNO}_3} = 534.1 \text{ g AgNO}_3$$

2. Nitrogen gas is reacted with hydrogen gas to form nitrogen trihydride.

a. Write and balance the chemical equation.



b. How many liters of nitrogen trihydride are produced at STP if 80.28 grams of hydrogen gas are reacted in an excess of nitrogen?

$$\frac{80.28 \text{ g H}_2}{1} \times \frac{1 \text{ mole H}_2}{2.016 \text{ g H}_2} \times \frac{2 \text{ moles NH}_3}{3 \text{ moles H}_2} \times \frac{22.4 \text{ L NH}_3}{1 \text{ mole NH}_3} = 594.7 \text{ L NH}_3$$

c. How many grams of hydrogen are needed to fully react 621.9 grams of nitrogen gas?

$$\frac{621.9 \text{ g N}_2}{1} \times \frac{1 \text{ mole N}_2}{28.014 \text{ g N}_2} \times \frac{3 \text{ moles H}_2}{1 \text{ mole N}_2} \times \frac{2.016 \text{ g H}_2}{1 \text{ mole H}_2} = 134.3 \text{ g H}_2$$

d. How many liters of nitrogen gas are needed to completely react 90.38 L of hydrogen gas at STP?

$$\frac{90.28 \text{ L H}_2}{1} \times \frac{1 \text{ mole H}_2}{22.4 \text{ L H}_2} \times \frac{1 \text{ mole N}_2}{3 \text{ mole H}_2} \times \frac{22.4 \text{ L N}_2}{1 \text{ mole N}_2} = 30.09 \text{ L N}_2$$

$$\frac{90.28 \text{ L H}_2}{1} \times \frac{\text{OR}}{3 \text{ L H}_2} = 30.09 \text{ L N}_2$$

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Table of Contents Stoichiometry 8b Extra Practice Problems Answers

1. Understanding the eBook Stoichiometry 8b Extra Practice Problems Answers
 - The Rise of Digital Reading Stoichiometry 8b Extra Practice Problems Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry 8b Extra Practice Problems Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry 8b Extra Practice Problems Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry 8b Extra Practice Problems Answers
 - Personalized Recommendations

- Stoichiometry 8b Extra Practice Problems Answers User Reviews and Ratings
- Stoichiometry 8b Extra Practice Problems Answers and Bestseller Lists
- 5. Accessing Stoichiometry 8b Extra Practice Problems Answers Free and Paid eBooks
 - Stoichiometry 8b Extra Practice Problems Answers Public Domain eBooks
 - Stoichiometry 8b Extra Practice Problems Answers eBook Subscription Services
 - Stoichiometry 8b Extra Practice Problems Answers Budget-Friendly Options
- 6. Navigating Stoichiometry 8b Extra Practice Problems Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Stoichiometry 8b Extra Practice Problems Answers Compatibility with Devices
 - Stoichiometry 8b Extra Practice Problems Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry 8b Extra Practice Problems Answers
 - Highlighting and Note-Taking Stoichiometry 8b Extra Practice Problems Answers
 - Interactive Elements Stoichiometry 8b Extra Practice Problems Answers
- 8. Staying Engaged with Stoichiometry 8b Extra Practice Problems Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry 8b Extra Practice Problems Answers
- 9. Balancing eBooks and Physical Books Stoichiometry 8b Extra Practice Problems Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry 8b Extra Practice Problems Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Stoichiometry 8b Extra Practice Problems Answers
 - Setting Reading Goals Stoichiometry 8b Extra Practice Problems Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Stoichiometry 8b Extra Practice Problems Answers
 - Fact-Checking eBook Content of Stoichiometry 8b Extra Practice Problems Answers

- Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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