

Name _____ Sex _____ Class _____

CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then use data for table. Show your calculations for questions 15–17 in the space below the table.



The molar masses of O_2 , NO , and NO_2 are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of O_2	Amount of NO	Amount of NO_2	Limiting Reactant	Amount of Excess Reactant
1 molecule	1 molecule	1 molecule	none	none
2 molecules	2 molecules	2 molecules	NO	1 molecule O_2
2 molecules	3 molecules	4 molecules	O_2	1 molecule NO
1.00 mol	2.00 mol	2.00 mol	NO	1.00 mol O_2
2.00 mol	4.00 mol	4.00 mol	NO	2.00 mol O_2
3.00 mol	6.00 mol	6.00 mol	O_2	3.00 mol NO
4.00 mol	8.00 mol	8.00 mol	O_2	4.00 mol NO
5.00 mol	10.00 mol	10.00 mol	O_2	5.00 mol NO
6.00 mol	12.00 mol	12.00 mol	O_2	6.00 mol NO
7.00 mol	14.00 mol	14.00 mol	O_2	7.00 mol NO
8.00 mol	16.00 mol	16.00 mol	O_2	8.00 mol NO
9.00 mol	18.00 mol	18.00 mol	O_2	9.00 mol NO
10.00 mol	20.00 mol	20.00 mol	O_2	10.00 mol NO
11.00 mol	22.00 mol	22.00 mol	O_2	11.00 mol NO
12.00 mol	24.00 mol	24.00 mol	O_2	12.00 mol NO

balanced equation mole ratio = 2 mol NO /1 mol O_2

32.00 g O_2 = 1 mol O_2 /32.00 g O_2 = 0.3125 mol O_2

20.00 g NO = 1 mol NO /30.01 g NO = 0.6664 mol NO

actual mole ratio = 0.6664 mol NO /0.3125 mol O_2 = 2.132 mol NO /1.000 mol O_2

Because the actual mole ratio of NO / O_2 is larger than the balanced equation mole ratio of NO / O_2 , there is an excess of NO / O_2 is the limiting reactant.

Mass of NO used = 0.3125 mol O_2 = 2 mol NO /1 mol O_2 = 0.6250 mol NO

0.6250 mol NO = 30.01 g NO /1 mol NO = 18.75 g NO

Mass of NO_2 produced = 0.6250 mol NO_2 = 46.01 g NO_2 /1 mol NO_2 = 28.75 g NO_2

Excess NO = 20.00 g NO - 18.75 g NO = 1.24 g NO

Name _____ Sex _____ Class _____

CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.4 Percent Yield

In your textbook, read about the yield of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of gallium nitride, a substance used in the manufacturing of some semiconductors for lasers.



In one experiment, the reaction yielded 7.42 g of the oxide from a 1.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and GaN are 69.72 g/mol and 122.62 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield volume = 7.42 g GaN
- mass of reactant = 1.00 g Ga
- number of moles of reactant = 1.00 g Ga \times $\frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}}$ = 7.00 g Ga = 1 mol Ga/69.72 g Ga
- number of moles of product = 0.00 mol Ga \times $\frac{2 \text{ mol GaN}}{2 \text{ mol Ga}}$ = 0.100 mol Ga = 2 mol GaN /19.72 mol Ga
- theoretical yield = 0.0000 mol GaN = $\frac{122.62 \text{ g GaN}}{1 \text{ mol GaN}}$ correct
- percent yield = $\frac{7.42 \text{ g GaN}}{122.62 \text{ g GaN}} \times 100$ = 7.42 g GaN /6.47 g GaN = 100

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1. Understanding the eBook Study Guide For Content Mastery Answer Key Chapter 12
 - The Rise of Digital Reading Study Guide For Content Mastery Answer Key Chapter 12
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide For Content Mastery Answer Key Chapter 12
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide For Content Mastery Answer Key Chapter 12
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide For Content Mastery Answer Key Chapter 12
 - Personalized Recommendations
 - Study Guide For Content Mastery Answer Key Chapter 12 User Reviews and Ratings
 - Study Guide For Content Mastery Answer Key Chapter 12 and Bestseller Lists
5. Accessing Study Guide For Content Mastery Answer Key Chapter 12 Free and Paid eBooks

- Study Guide For Content Mastery Answer Key Chapter 12 Public Domain eBooks
 - Study Guide For Content Mastery Answer Key Chapter 12 eBook Subscription Services
 - Study Guide For Content Mastery Answer Key Chapter 12 Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Study Guide For Content Mastery Answer Key Chapter 12 Compatibility with Devices
 - Study Guide For Content Mastery Answer Key Chapter 12 Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide For Content Mastery Answer Key Chapter 12
 - Highlighting and Note-Taking Study Guide For Content Mastery Answer Key Chapter 12
 - Interactive Elements Study Guide For Content Mastery Answer Key Chapter 12
 8. Staying Engaged with Study Guide For Content Mastery Answer Key Chapter 12
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide For Content Mastery Answer Key Chapter 12
 9. Balancing eBooks and Physical Books Study Guide For Content Mastery Answer Key Chapter 12
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide For Content Mastery Answer Key Chapter 12
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Study Guide For Content Mastery Answer Key Chapter 12
 - Setting Reading Goals Study Guide For Content Mastery Answer Key Chapter 12
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Study Guide For Content Mastery Answer Key Chapter 12
 - Fact-Checking eBook Content of Study Guide For Content Mastery Answer Key Chapter 12
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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