

For questions 1 - 7, choose true or false.

- The rate of a given reaction is constant at any specified temperature. **False**
- A balanced equation does not indicate the mechanism by which a reaction is proceeding. **True**
- The rate constant is the rate of reaction at constant temperature. **False**
- In the reaction  $\text{H}_2 + \text{I}_2 \leftrightarrow 2 \text{HI}$ , only two molecules ( $\text{H}_2$  and  $\text{I}_2$ ) react so this reaction is a bimolecular reaction. **False**
- Relatively few reaction mechanisms have been completely characterized. **True**  
(Special ingenuity and techniques are needed to detect fast steps in a reaction mechanism.)
- The rate-determining step of a reaction is the step that proceeds fastest. **False**
- The half-life of a reaction is the time required for half of the original concentration of the limiting reagent to be consumed in the reaction. **True**
- Factors that determine reaction rates include all of the following except:
  - size of reactant particles
  - magnitude of the equilibrium constant** (Large values of  $k$  do not imply large rates.)
  - reaction temperature
  - concentration of reactants
- For the reaction  $2\text{A} + \text{B} \rightarrow \text{C}$ , the reaction rate is:
  - $R = k[\text{A}]^2 [\text{B}]$
  - $R = k[\text{A}] [\text{B}]$
  - $R = k[\text{A}]^2$
  - possibly not any of the above** (A valid reaction rate expression usually cannot be written from the overall reaction.)
- For a reaction of the type  $\text{A} + \text{B} \rightarrow 2\text{C}$ , it is found that doubling the amount of A causes the reaction rate to be four times as great but that doubling the amount of B has no apparent effect on the rate. The rate equation that is most correct is:
  - $R = k[\text{A}]^2$  (second order in A and zero order in B)**
  - $R = k[\text{A}] [\text{B}]$
  - $R = k[\text{A}]$
  - $R = \frac{k[\text{C}]^2}{[\text{A}] [\text{B}]}$
- Which of the following is the rate equation for a termolecular elementary reaction?
  - $R = k [\text{A}]$
  - $R = k [\text{A}]^2$
  - $R = k [\text{A}] [\text{B}]$
  - $R = k [\text{A}] [\text{B}] [\text{C}]$  (three component collision)**

# Ap Chemistry Chapter 12 Test

**Terry C. Jones**



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