

STUDY GUIDE FOR LECTURE EXAM (P2) (CHAPTERS 4-7)

1. Chapter 4

1. Define the types of chemical bonds and their properties

Electrostatic interactions

- Ionic bonds: electrostatic attraction
- Dipole-dipole interactions: partial charges
- Hydrogen bonds: partial charges

Covalent bonds: electron sharing

- Polar covalent bonds: unequal sharing of electrons due to electronegativity differences (more than 1.7 is ionic)

2. For each bond, describe the bond parameters, bond strength, and bond length

Electrostatic interactions: ionic, dipole-dipole, hydrogen

Covalent bonds: single, double

Electrostatic interactions: ionic, dipole-dipole, hydrogen

Covalent bonds: single, double, triple

3. Describe the structure, properties, strength, and electronegativity of the atoms

Electrostatic interactions: ionic, dipole-dipole, hydrogen

Covalent bonds: single, double, triple

Electrostatic interactions: ionic, dipole-dipole, hydrogen and covalent

4. List the elements of the periodic table

Electrostatic interactions

- Ionic bonds: ionic, dipole-dipole, hydrogen
- Dipole-dipole interactions: partial charges
- Hydrogen bonds: partial charges
- Polar covalent bonds: unequal sharing of electrons (more than 1.7 is ionic)
- Nonpolar covalent bonds: equal sharing of electrons
- Single, double, triple
- Electronegativity: ability to attract electrons

5. List all the elements of the periodic table

Electrostatic interactions: ionic, dipole-dipole, hydrogen and covalent

Covalent bonds: single, double, triple

Electrostatic interactions: ionic, dipole-dipole, hydrogen and covalent

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