



# COVALENT BONDING

## Practice Problems

Do your answers to the following problems.

### SECTION 5.1 MOLECULAR COMPOUNDS

- Classify each of the following as either an ionic or a molecular compound.
  - $\text{NaCl}$
  - $\text{H}_2\text{O}$
  - $\text{HCl}$
  - $\text{KNO}_3$
  - $\text{C}_2\text{H}_6$
  - $\text{H}_2\text{SO}_4$
  - $\text{CaCl}_2$
  - $\text{C}_2\text{H}_5\text{OH}$
  - $\text{CO}_2$
  - $\text{H}_2\text{O}_2$
  - $\text{CH}_4$
  - $\text{H}_2\text{S}$

- Classify each of the following as either an ionic or a molecular compound.
  - $\text{K}_2\text{SO}_4$
  - $\text{C}_2\text{H}_6$
  - $\text{H}_2\text{O}$
  - $\text{HCl}$
  - $\text{CaCl}_2$
  - $\text{C}_2\text{H}_5\text{OH}$
  - $\text{CO}_2$
  - $\text{H}_2\text{O}_2$
  - $\text{CH}_4$
  - $\text{H}_2\text{S}$
- Write the formula for each of the following molecular compounds.
  - Carbon dioxide
  - Hydrogen peroxide
  - Hydrogen cyanide
  - Hydrogen sulfide
  - Hydrogen chloride
  - Hydrogen bromide
  - Hydrogen iodide
  - Hydrogen fluoride
  - Hydrogen selenide
  - Hydrogen telluride
  - Hydrogen phosphide
  - Hydrogen arsenide
  - Hydrogen stibide
  - Hydrogen selenide
  - Hydrogen telluride
  - Hydrogen phosphide
  - Hydrogen arsenide
  - Hydrogen stibide
  - Hydrogen selenide
  - Hydrogen telluride
  - Hydrogen phosphide
  - Hydrogen arsenide
  - Hydrogen stibide

### SECTION 5.2 THE NATURE OF COVALENT BONDING

- Draw the Lewis structure for hydrogen fluoride.  $\text{H}-\text{F}$
- Draw the Lewis structure for hydrogen chloride.  $\text{H}-\text{Cl}$
- Draw the Lewis structure for hydrogen bromide.  $\text{H}-\text{Br}$
- Draw the Lewis structure for hydrogen iodide.  $\text{H}-\text{I}$
- Draw the Lewis structure for hydrogen cyanide.  $\text{H}-\text{C}\equiv\text{N}$
- Draw the Lewis structure for hydrogen sulfide.  $\text{H}-\text{S}-\text{H}$
- Draw the Lewis structure for hydrogen selenide.  $\text{H}-\text{Se}-\text{H}$
- Draw the Lewis structure for hydrogen telluride.  $\text{H}-\text{Te}-\text{H}$
- Draw the Lewis structure for hydrogen phosphide.  $\text{H}-\text{P}-\text{H}_3$
- Draw the Lewis structure for hydrogen arsenide.  $\text{H}-\text{As}-\text{H}_3$
- Draw the Lewis structure for hydrogen stibide.  $\text{H}-\text{Sb}-\text{H}_3$
- Draw the Lewis structure for hydrogen selenide.  $\text{H}-\text{Se}-\text{H}_2$
- Draw the Lewis structure for hydrogen telluride.  $\text{H}-\text{Te}-\text{H}_2$
- Draw the Lewis structure for hydrogen phosphide.  $\text{H}-\text{P}-\text{H}_3$
- Draw the Lewis structure for hydrogen arsenide.  $\text{H}-\text{As}-\text{H}_3$
- Draw the Lewis structure for hydrogen stibide.  $\text{H}-\text{Sb}-\text{H}_3$



### SECTION 5.3 BONDING: THE ORBITALS

- Which orbital and suborbital are represented by the following?
  - $2s$
  - $2p$
  - $3s$
  - $3p$
  - $4s$
  - $4p$
  - $5s$
  - $5p$
  - $6s$
  - $6p$
  - $7s$
  - $7p$
- Which orbital and suborbital are represented by the following?
  - $2s$
  - $2p$
  - $3s$
  - $3p$
  - $4s$
  - $4p$
  - $5s$
  - $5p$
  - $6s$
  - $6p$
  - $7s$
  - $7p$

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