

CHEMISTRY REVIEW: KEY CHAPTER 9 – CHEMICAL NAMES AND FORMULAS

- In a Stock name such as iron (II) sulfate, the roman numeral tells us **C**
 - How many atoms of Fe are in one formula unit
 - How many sulfate ions can be attached to the iron atom
 - The charge on each Fe ion
 - The total positive charge of the formula unit
- The result of changing a subscript in a correctly written chemical formula is to **c**
 - Change the number of moles represented in the formula
 - Change the charges on the other ions in the compound
 - Change the formula so that it no longer represents the compound
 - Have no effect on the formula
- Derive a generalization for when an acid name will end in the suffix **-ic** or **-ous**
If ending is -ic, the anion ended in -ate; if ending is -ous, the anion ended in -ite
- Derive a generalization for when an acid name will begin with the prefix **hydro-**
If prefix is hydro-, the anion ended in -ide. This is most likely a monatomic ion.
- Fill in the blanks in the table below

Compound Name	Formula
Aluminum sulfide	Al_2S_3
Aluminum sulfate	$Al_2(SO_4)_3$
Lead (II) chloride	$PbCl_2$
Ammonium phosphate	$(NH_4)_3PO_4$
Hydrobromic acid	HBr
Dinitrogen pentoxide	N_2O_5
Copper (I) carbonate	Cu_2CO_3
Sodium sulfide	Na_2S
Ammonium phosphate	$(NH_4)_3PO_4$
Tin (IV) sulfide	$Sn(SO_4)_2$
Nitrous acid	HNO_2
Magnesium perchlorate	$Mg(ClO_4)_2$
Iron (II) nitrate	$Fe(NO_3)_2$
Iron (III) nitrate	$Fe(NO_3)_3$
Cobalt (II) oxide	CoO
Sulfur trioxide	SO_3
Boron pentafluoride	BF_5
Tetraphosphorus octasulfide	P_4S_8
Tetraoxide nonoxide	TeO_9

- Name the following compound and identify as molecular, ionic or acid
 - $Ca(HCO_3)_2$ calcium bicarbonate - **I**
 - $Fe(ClO_2)_3$ iron (III) chlorite - **I**
 - $Fe(ClO_4)_2$ iron (II) chlorate - **I**
 - N_2O_5 dinitrogen pentoxide - **M**

Chapter 9 Chemical Names And Formulas Answers

Sharon J. Sherman



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oxidation $\text{CO} + \text{CO}_2 + 2\text{e}^-$ reduction $\text{Fe} + 2\text{e}^- + 2\text{H}^+$ $24 \text{ 3 Ag} + \text{Ni}^{3+} \rightarrow 3 \text{ Ag} + \text{Ni}$ oxidation reduction $25 \text{ Cu} + \text{Pb}^{2+} \rightarrow \text{Cu}^{2+} + \text{Pb}$ oxidation reduction 26 O_2

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 $5\text{CO} + \text{I}_2 \rightarrow 5\text{CO}_2 + \text{I}_2$ $\text{b} \text{ O}_5 + 5\text{CO} \rightarrow \text{I}_2 + 5\text{CO}_2$

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