

Laboratory Report – Periodic Table

Type Name Here:

Part A

1. Place 9 clean small test tubes in a test tube rack. The test tubes should be clean but do not have to be dry because aqueous solutions will be used.
2. To each test tube add 5 drops of the solution (0.1 M) indicated in the data table.
3. Next add 5 drops of silver nitrate (0.5 M) to each of the 9 test tubes. If there is no apparent change be sure to gently swirl the test tube to mix the reagents (a stirring rod can be used). If the contents of the test tube are cloudy, a solid precipitate was formed. If the contents of the test tube are clear, no reaction occurred. Record the results in the data table writing "ppt" and their color for precipitates and "NR" for no reaction.
4. Your instructor will inform you of where to dispose of the chemicals used. Rinse the test tubes and repeat the procedure two more times first using ammonium carbonate (0.5 M) in step 3 and then again using sodium phosphate (0.5 M) in step 3.

Test Tube	Solution	Test Ion	Ag ⁺	CO ₃ ²⁻	PO ₄ ³⁻
1	Barium nitrate	Ba ²⁺		White ppt	White ppt
2	Ammonium chloride	Cl ⁻	White ppt		
3	Calcium nitrate	Ca ²⁺		White ppt	White ppt
4	Lithium nitrate	Li ⁺			
5	Ammonium iodide	I ⁻	Yellowish ppt		
6	Potassium nitrate	K ⁺			
7	Sodium nitrate	Na ⁺			
8	Strontium nitrate	Sr ²⁺		White ppt	White ppt
9	Ammonium bromide	Br ⁻	Off-White ppt		

Chem Fax Lab Answers

Ying-Ying Zheng



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Table of Contents Chem Fax Lab Answers

1. Understanding the eBook Chem Fax Lab Answers
 - The Rise of Digital Reading Chem Fax Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chem Fax Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chem Fax Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chem Fax Lab Answers
 - Personalized Recommendations
 - Chem Fax Lab Answers User Reviews and Ratings
 - Chem Fax Lab Answers and Bestseller Lists
5. Accessing Chem Fax Lab Answers Free and Paid eBooks
 - Chem Fax Lab Answers Public Domain eBooks
 - Chem Fax Lab Answers eBook Subscription Services
 - Chem Fax Lab Answers Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Chem Fax Lab Answers Compatibility with Devices
 - Chem Fax Lab Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chem Fax Lab Answers
 - Highlighting and Note-Taking Chem Fax Lab Answers
 - Interactive Elements Chem Fax Lab Answers
8. Staying Engaged with Chem Fax Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chem Fax Lab Answers
9. Balancing eBooks and Physical Books Chem Fax Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chem Fax Lab Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chem Fax Lab Answers
 - Setting Reading Goals Chem Fax Lab Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chem Fax Lab Answers
 - Fact-Checking eBook Content of Chem Fax Lab Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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