

## Test "Chemistry"

## Chapter # 3

### Objective part

Time: 15 min

Total marks: 12

Q.1- Every question has four possible answers. Encircle the correct option.

- i) The atomic radii of the elements in periodic table;  
a) increase from left to right in a period  
b) increase from top to bottom in a group  
c) do not change from left to right in a period  
d) decrease from top to bottom in a group
- ii) Mendeleev Periodic table was based upon the:  
a) electronic configuration  
b) atomic mass  
c) atomic number  
d) electro negativity
- iii) The amount of energy given out when an electron is added to an atom is called;  
a) ionization energy  
b) lattice energy  
c) electron affinity  
d) electro negativity
- iv) Long form of periodic table is constructed on the basis of:  
a) Mendeleev postulate  
b) atomic number  
c) atomic mass  
d) mass number
- v) Along the period which of the following decreases:  
a) electro negativity  
b) ionization energy  
c) atomic radius  
d) electron affinity
- vi) Mark the incorrect statement about the ionization energy:  
a) it is measured  $\text{kJmol}^{-1}$   
b) it is the absorption of energy  
c) it decreases in a period  
d) it decreases in groups
- vii) Transition elements are:  
a) all gases  
b) all metals  
c) all non metals  
d) all metalloids
- viii) S-f series of inner transition elements are called:  
a) lanthanides  
b) actinides  
c) halogens  
d) alkali metals
- ix) If we move from left to right in a period the value of ionization energy:  
a) remain same  
b) decreases  
c) increases  
d) not effected
- x) Which scientist has given the idea of actives in periodic table:  
a) Mendeleev  
b) Al-razi  
c) Newland  
d) dobernier
- xi) Shielding effect is greater in atoms with greater number of:  
a) proton  
b) electron  
c) neutron  
d) positron
- xii) Which is the incomplete period in the periodic table:  
a) 4th period  
b) 2nd period  
c) 7th period  
d) 1st period

# Chemistry Chapter 3 Test Answers

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