

Chapter 12: Liquids, Solids, and Intermolecular Forces

Why are some substances solids, while others are liquids or gases at room temperature and pressure (25°C, 1 atm)?

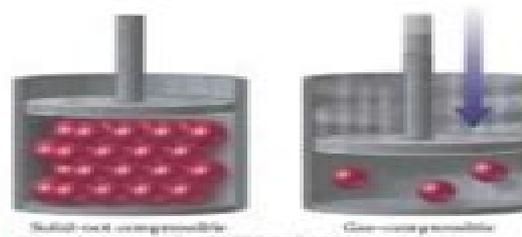
The different physical states of matter (SOLID, LIQUID, GAS) exist primarily because of differences in **INTERMOLECULAR FORCES** that act between particles of these substances.

Substances that are **GASES** at room temp. and pressure:

- Have **very weak** intermolecular interactions
- Have particles that are **far apart**
 - The particles have complete freedom from each other
- Have **low density** compared to liquids and solids
- Expand to fill their container
 - The particles are constantly flying around, bumping into each other and their container
- Can be **easily compressed** into a smaller volume
- **Are fluid** - exhibits a smooth, continuous flow as it moves
- Are either **monoatomic atoms** (He, Ne) or are **molecules with covalent bonding**



Gas



Chemistry Chapter 12 Review Liquids And Solids Answers

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