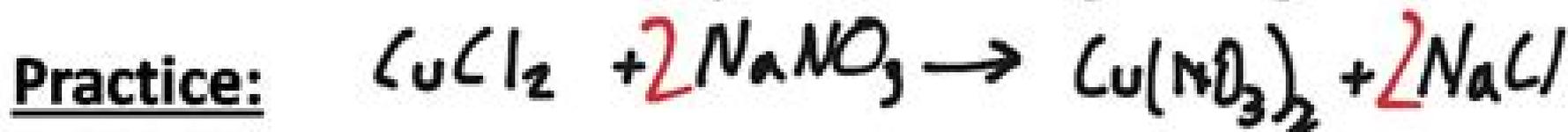


9.2 – Stoichiometry – Limiting Reagent



20 grams of copper (II) chloride react with 25 grams of sodium nitrate. Identify the limiting reagent.

$$20\text{g } \cancel{\text{CuCl}_2} \times \frac{1 \text{ mol } \cancel{\text{CuCl}_2}}{137.452\text{g } \cancel{\text{CuCl}_2}} = 0.149 \text{ mol } \cancel{\text{CuCl}_2} \times \frac{2 \text{ mol } \text{NaNO}_3}{1 \text{ mol } \cancel{\text{CuCl}_2}} = 0.298 \text{ mol } \text{NaNO}_3$$

HAVE Need

$$25\text{g } \cancel{\text{NaNO}_3} \times \frac{1 \text{ mol } \cancel{\text{NaNO}_3}}{84.994\text{g } \cancel{\text{NaNO}_3}} = 0.294 \text{ mol } \cancel{\text{NaNO}_3} \times \frac{1 \text{ mol } \text{CuCl}_2}{2 \text{ mol } \cancel{\text{NaNO}_3}} = 0.147 \text{ mol } \text{CuCl}_2$$

You need 0.298 mol NaNO_3 , but you have 0.294 mol NaNO_3

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B I. Bhatt



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Stoichiometry Unit Project Luann Marie Decker, 1998 **Stoichiometry** Sydney Young, 1918 This introduction was originally prefixed to Dr Findlay's Phase rule which was the first volume of the series issued. It belongs properly however to this volume and is therefore included here. Note p vii Introduction issued also separately in 1904 **Formulation and Stoichiometry** Emil J. Margolis, 2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry. As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all. The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities. For the serious student this means more often than not the frustrations of questions unanswered. The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer. Distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media. Quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions. Intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction **Formulation and Stoichiometry** Emil J Margolis, Stoichiometry B. I. Bhatt, Shuchen B. Thakore, 2010 **Stoichiometry** B I. Bhatt, 2010 *Stoichiometry* B. I. Bhatt, 1996 Stoichiometry, 4E B. I. Bhatt, 2004-03

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