

QUESTION 1

Volume of pipette used 25.00 cm^3
 (V₀)

Titration	Rough	1	2	3
Final burette reading (cm ³)	20.80	41.40	20.40	40.90
Initial burette reading (cm ³)	0.00	20.80	0.00	20.40
Volume of A used	20.80	20.60	20.40	20.50

② ^{average} Volume of A used = $\frac{20.60 + 20.40 + 20.50}{3}$
 $= 20.50 \text{ cm}^3$

③(i) Concentration of A in mol/dm³ = $\frac{\text{conc. in g/dm}^3}{\text{molar mass}}$
 $= \frac{4.4 \text{ g/dm}^3}{2[13] + 32 + 4[16]}$
 $= \frac{4.4}{98} = 0.0449 \text{ mol/dm}^3$

③(ii) Concentration of B in mol/dm³, C_B

Using $\frac{C_A V_A}{C_B V_B} = \frac{n_A}{n_B}$

$$\frac{0.0449 \times 20.50}{C_B \times 25} = \frac{1}{1}$$

$$C_B = \frac{0.0449 \times 20.50}{25} = 0.03287 \text{ mol/dm}^3$$

Chemistry Practical Question And Answer

Oswaal Editorial Board



Chemistry Practical Question And Answer:

Chemistry Practical Question And Answer Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has are more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Chemistry Practical Question And Answer**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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Table of Contents Chemistry Practical Question And Answer

1. Understanding the eBook Chemistry Practical Question And Answer
 - The Rise of Digital Reading Chemistry Practical Question And Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Practical Question And Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Practical Question And Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Practical Question And Answer
 - Personalized Recommendations
 - Chemistry Practical Question And Answer User Reviews and Ratings

- Chemistry Practical Question And Answer and Bestseller Lists
- 5. Accessing Chemistry Practical Question And Answer Free and Paid eBooks
 - Chemistry Practical Question And Answer Public Domain eBooks
 - Chemistry Practical Question And Answer eBook Subscription Services
 - Chemistry Practical Question And Answer Budget-Friendly Options
- 6. Navigating Chemistry Practical Question And Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Practical Question And Answer Compatibility with Devices
 - Chemistry Practical Question And Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Practical Question And Answer
 - Highlighting and Note-Taking Chemistry Practical Question And Answer
 - Interactive Elements Chemistry Practical Question And Answer
- 8. Staying Engaged with Chemistry Practical Question And Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Practical Question And Answer
- 9. Balancing eBooks and Physical Books Chemistry Practical Question And Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Practical Question And Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Practical Question And Answer
 - Setting Reading Goals Chemistry Practical Question And Answer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Practical Question And Answer
 - Fact-Checking eBook Content of Chemistry Practical Question And Answer
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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